



INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(51) International Patent Classification 5 : C09D 5/08	A1	(11) International Publication Number: WO 93/15154 (43) International Publication Date: 5 August 1993 (05.08.93)
(21) International Application Number: PCT/US93/00137 (22) International Filing Date: 15 January 1993 (15.01.93) (30) Priority data: 4/40571 31 January 1992 (31.01.92) JP (71) Applicant (for all designated States except US): HENKEL CORPORATION [US/US]; 140 Germantown Pike, Suite 150, Plymouth Meeting, PA 19462 (US). (72) Inventors; and (75) Inventors/Applicants (for US only) : HONDA, Takumi [JP/JP]; 25-4-303, Daikan-cho, Hiratsuka-shi, Kanagawa-ken 254 (JP). NAITO, Kazuhisa [JP/JP]; 564-10, Kanagawa, Chigasaki-shi, Kanagawa-ken 253 (JP).		(74) Agent: ORTIZ, Daniel, S.; Henkel Corporation, 140 Germantown Pike, Suite 150, Plymouth Meeting, PA 19462 (US). (81) Designated States: AU, BR, CA, US, European patent (AT, BE, CH, DE, DK, ES, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE). Published <i>With international search report.</i>
(54) Title: AUTODEPOSITION COATING COMPOSITION (57) Abstract An aqueous autodeposition coating composition that is essentially free of chromium compounds and which produces a film with a much better adherence and corrosion resistance than heretofore obtained, but which does so without a rinse treatment with a chromium-containing solution prior to film curing. The aqueous autodeposition coating composition has a pH of about 1.6 to about 5 and contains (i) water-dispersable or water-soluble organic film-forming resin; (ii) at least one of fluoride ions and fluoride ions and complex fluoride ions; (iii) ions of at least one metal selected from zinc, cobalt, manganese, nickel, iron, and aluminum; and (iv) tungstate ion and/or molybdate ion.		

FOR THE PURPOSES OF INFORMATION ONLY

Codes used to identify States party to the PCT on the front pages of pamphlets publishing international applications under the PCT.

AT	Austria	FR	France	MR	Mauritania
AU	Australia	GA	Gabon	MW	Malawi
BB	Barbados	GB	United Kingdom	NL	Netherlands
BE	Belgium	GN	Guinea	NO	Norway
BF	Burkina Faso	GR	Greece	NZ	New Zealand
BG	Bulgaria	HU	Hungary	PL	Poland
BJ	Benin	IE	Ireland	PT	Portugal
BR	Brazil	IT	Italy	RO	Romania
CA	Canada	JP	Japan	RU	Russian Federation
CF	Central African Republic	KP	Democratic People's Republic of Korea	SD	Sudan
CG	Congo	KR	Republic of Korea	SE	Sweden
CH	Switzerland	KZ	Kazakhstan	SK	Slovak Republic
CI	Côte d'Ivoire	LJ	Liechtenstein	SN	Senegal
CM	Cameroon	LK	Sri Lanka	SU	Soviet Union
CS	Czechoslovakia	LJ	Luxembourg	TD	Chad
CZ	Czech Republic	MC	Monaco	TG	Togo
DE	Germany	MG	Madagascar	UA	Ukraine
DK	Denmark	ML	Mali	US	United States of America
ES	Spain	MN	Mongolia	VN	Viet Nam
FI	Finland				

AUTODEPOSITION COATING COMPOSITION

FIELD OF THE INVENTION

The invention relates to an aqueous autodeposition coating composition in the form of an acidic coating composition (pH approximately 1.6 to 5) that contains a water-dispersible or water-soluble organic film-forming resin; fluoride ions or fluoride ions and complex fluoride ions; one or more ions selected from ions of the following metals: zinc, cobalt, manganese, nickel, iron, and aluminum; and tungstate ion and/or molybdate ion. The aqueous autodeposition coating composition is capable of forming a highly corrosion-resistant, strongly adherent resin film on metal surfaces when brought into contact with the surface of a metal, for example, a ferriferous metal, zinciferous metal, aluminiferous metal, magnesium-based metal, and the like.

RELATED ART

Autodeposition coating compositions are acidic coating compositions that contain an organic film-forming resin and form a resin film on a metal surface when brought into contact with the metal surface. Autodeposition coating compositions are disclosed in, for example,

Japanese Patent Publication Number Sho 47-17630
[17,630/1972],
Japanese Patent Publication Number Sho 48-14412
[14,412/1973],
Japanese Patent Publication Number Sho 52-21006
[21,006/1977],
Japanese Patent Publication Number Sho 52-35692
[35,692/1977],
Japanese Patent Publication Number Sho 53-15093
[15,093/1978],
Japanese Patent Publication Number Sho 53-44949
[44,949/1978],
Japanese Patent Application Laid Open [Kokai or
Unexamined] Number Sho 60-58474 [58,474/1985],
Japanese Patent Application Laid Open Number Sho
61-168673 [168,673/1986], and
Japanese Patent Application Laid Open Number Sho
61-246267 [246,267/1986].

A characteristic feature of the known coating compositions is their ability to lay down a resin film upon immersion of a clean metal surface in the particular

coating composition. Moreover, the coating film's thickness and weight increase with immersion time. In the case of these compositions, the chemical activity of the coating composition overlying the metal surface (metal ion eluted from the metal surface by etching induces association of the resin particles with resulting deposition on the metal surface) results in the efficient deposition of a resin film on the metal surface without the application, as in electrodeposition, of an external electrical source.

One example of a metal ion-containing coating composition is disclosed in Japanese Patent Publication Number Sho 52-35692. This publication discloses the use of a composition that contains a ferric ion-supplying iron compound to coat ferriferous metal. A coating composition that contains various metal ions (not including the ferric ion) is disclosed in Japanese Patent Publication Number Sho 53-44949 for the coating of various types of metal substrates. However, the demands made on coating performance have recently become more severe, particularly with regard to adherence and corrosion resistance. As disclosed in United States Patent 3,647,567, United States Patent 4,030,945, Japanese Patent Publication Number Sho 53-15093, and Japanese Patent Application Laid Open Number Sho 61-168673, a rinse treatment with a chromium-containing solution or aqueous alkali solution, etc. (post-treatment directly after autodeposition) may be implemented prior to film curing (prior to baking/drying) in order thereby to

provide the ultimately produced film with better corrosion resistance and better adherence. In this case, the presence of a chromium compound is particularly effective for enhancing film performance and particularly the corrosion resistance provided by the film.

However, environmental considerations provide a strong incentive against the use of chromium compounds, for example, a chromium rinse, and of late there has been strong demand for the use of chromium-free coating compositions for the coating of metals and the like. In addition, the ability to provide a film with excellent adherence and corrosion resistance with fewer process steps (contraction) has also become an important consideration. All in all, it is an important problem to improve the adherence and corrosion resistance of the final coating without using chromium compounds. Nevertheless, films generated by prior chromium-free coating compositions have not been fully satisfactory in their critical properties (high adherence, high corrosion resistance).

BRIEF DESCRIPTION OF THE INVENTION

The present invention provides an essentially chromium compound-free film which has much better adherence and corrosion resistance than the films formed by prior-art coating compositions, but which is produced without a rinse treatment (for example, with a chromium-containing solution) prior to film curing.

The aqueous autodeposition coating composition of the invention has a pH of about 1.6 to about 5 and contains

- (i) water-dispersible or water-soluble organic film-forming resin;
- (ii) at least one of fluoride ions or fluoride ions and complex fluoride ions;
- 5 (iii) additional metal ions of at least one metal selected from the group consisting of zinc, cobalt, manganese, nickel, iron, and aluminum; and
- (iv) at least one of tungstate ion and molybdate ion.

A highly adherent, strongly corrosion-resistant resin
10 film is formed on the metal surface when the aqueous auto-deposition coating composition of the present invention is brought into contact with a metal surface, for example, the surface of ferriferous, zinciferous, aluminiferous, or magnesium-based metal.

15 The composition of the invention provides films with a better adherence and corrosion resistance than the films afforded by prior autodeposition coating compositions due to an acidic coating composition (pH about 1.6 to about 5) that contains organic film-forming resin; fluoride ion or
20 fluoride ion and complex fluoride ion; one or more additional metal ions selected from the group consisting of zinc, cobalt, manganese, nickel, iron, and aluminum; and also tungstate ion and/or molybdate ion.

In addition, the present invention provides a highly
25 adherent, highly corrosion-resistant film without having to carry out a post-treatment (precure rinse treatment, for example, with a chromium-containing solution) that have heretofore been implemented for the purpose of improving

the adherence and corrosion resistance of the films produced by autodepositing coating compositions.

DETAILED DESCRIPTION OF THE INVENTION

Resin of the type disclosed in Japanese Patent Application Sho 61-168673 is an example of resin highly suitable for use as the organic film-forming resin in the present invention.

The organic film-forming resin useful in the practice of the present invention is exemplified by the following: urethane resins, epoxy resins, polyester resins, and polymer resins composed of one or more monomers selected from methyl acrylate, ethyl acrylate, n-butyl acrylate, 2-hydroxyethyl acrylate, 2-hydroxypropyl acrylate, 2-ethylhexyl acrylate, methyl methacrylate, ethyl methacrylate, n-butyl methacrylate, 2-hydroxyethyl methacrylate, 2-hydroxypropyl methacrylate, glycidyl acrylate, glycidyl methacrylate, acrylamide, methacrylamide, acrylonitrile, ethylene, styrene, vinyl chloride, vinylidene chloride, vinyl acetate, acrylic acid, methacrylic acid, and the like.

The organic film-forming resin useful in the practice of the invention may be anionic, cationic, nonionic, or amphoteric and is not specifically restricted in this regard. The content of resin solids in the coating composition preferably falls within the range of 5 to 550 g/L and more preferably falls within the range of 50 to 100 g/L.

Sources for the fluoride ion and complex fluoride ion

useful in the practice of the invention are, for example, zirconium hydrogen fluoride, titanium hydrogen fluoride, silicon hydrogen fluoride, boron hydrogen fluoride, hydrofluoric acid, and the ammonium, lithium, sodium, and potassium salts of the preceding acids. The content of fluoride or fluoride ion and complex fluoride ion in the coating composition preferably falls within the range of 0.1 to 5 g/L as fluorine and more preferably falls within the range of 0.5 to 3 g/L as fluorine.

The pH of the coating composition of the invention should be maintained within the range of about 1.6 to about 5. Formation of the resin film becomes problematic when the pH is substantially outside this range. The pH of the coating composition may be regulated using one or more acids selected from inorganic acids such as the acids listed above as sources of fluoride and complex fluoride ion and their salts, as well as nitric acid, phosphoric acid, and boric acid; and organic acids selected from phytic acid and tannic acid. The addition of nitric acid, phosphoric acid, boric acid, phytic acid, or tannic acid has the effect of improving film adherence to the substrate.

The zinc, cobalt, manganese, nickel, iron, and aluminum metal ions useful in the practice of the invention (hereinafter additional metal ions) can be supplied using the carbonates, nitrates, phosphates, sulfates, hydroxides, oxides, and chlorides of the corresponding metals. The metal ion is taken into the resin film during film

formation (deposition process), and it functions as a crosslinker for the resin during the heating carried out during film drying and curing. The metal ion bonds with the functional groups in the resin to bring about a substantial increase in the molecular weight of the resin and thereby improve the corrosion resistance of the film formed on the metal. The additional metal ions are present in solution in the composition in the range of from about 0.1 to about 20 gram/liter preferably from about 0.2 to about 10 grams/liter and more preferably from about 0.3 to about 6 grams/liter.

A critical feature of the present invention is the additional presence of at least one of tungstate ion and molybdate ion in a coating composition that contains organic film-forming resin; fluoride and/or complex fluoride ion; and one or more ions selected from the ions of the following metals: zinc, cobalt, manganese, nickel, iron, and aluminum.

Applicant will not be bound by theory but Applicant's theory is that the tungstate ion and/or molybdate ion present in the aqueous autodeposition coating composition of the invention forms a complex with a portion of the aforementioned metal ions. The metal (ion) in the complex appears to substitute for metal ion eluted from the surface of the metal workpiece, which results in efficient deposition onto the surface of the metal workpiece and ultimately coverage of the metal surface together with the organic film-forming resin.

In the case of the prior-art chromium rinses, the chromium (ion) taken into the resin film is largely present relatively near the surface of the resin film. In contrast to this, due to the presence of tungstate ion and/or molybdate ion in the coating composition, the metal ion taken into the resin film is largely present not only within the resin film, but also in particular in the vicinity of the metal surface of the substrate to a much greater degree than for prior-art coating compositions. This results in the production of a film having excellent adherence and corrosion resistance.

The tungstate ion used in the present invention can be supplied in the form of tungstic acid, sodium tungstate, calcium tungstate, potassium tungstate, and the like. The molybdate ion used in the present invention can be supplied in the form of molybdic acid, sodium molybdate, calcium molybdate, potassium molybdate, and the like. The content of tungstate ion and molybdate ion in the coating composition should be 0.1 to 5 g/L and is preferably 0.3 to 2 g/L.

The coating composition of the present invention may also contain an oxidant as an optional material. In addition, the coating composition of the present invention can contain pigment to impart color to the film.

Examples of the present invention and comparison examples are presented below.

Composition Production Examples A through Q

Aqueous coating compositions were prepared using an

acrylic emulsion (41.5% solids, Rhoplex WL-91 from the Rohm & Haas Company) and the other components as reported in Table 1 (brought to a total of 1 L with de-ionized water).

Examples 1 through 16

5 The aqueous coating compositions prepared in Composition Production Examples A through P were used in these examples. The aqueous coating composition baths were held at approximately 20°C to 22°C. Preliminarily cleaned cold-rolled steel sheet, galvanized hot-dipped
10 zinc-plated steel sheet, electrogalvanized steel sheet, or aluminum sheet (size = 70 x 150 x 1 mm in all cases) was treated in the particular example by immersion for 180 seconds. After a water rinse, the sample was dried in an oven for 20 minutes at 180°C and was then submitted to the
15 various tests. The results of the performance testing of the test sheets are reported in Table 2.

Comparison Examples 1 through 3

 The aqueous coating composition prepared in Composition Production Example Q was used in the comparison
20 examples. The aqueous coating composition baths were held at approximately 20°C to 22°C. Preliminarily cleaned cold-rolled steel sheet, galvanized hot-dipped zinc-plated steel sheet, or electrogalvanized steel sheet (size = 70 x 150 x 1 mm in all cases) were treated by
25 immersion for 180 seconds. After a water rinse, the sample was dried in an oven for 20 minutes at 180°C and was then submitted to the various tests. The results of the performance testing of the test sheets are reported in

Table 2.

Comparison Example 4

The aqueous coating composition prepared in Composition Production Example Q was used in this comparison example. The aqueous coating composition bath was held at approximately 20°C to 22°C. Preliminarily cleaned cold-rolled steel sheet (size = 70 x 150 x 1 mm) was treated by immersion for 180 seconds. After a water rinse, the sample was immersed in an aqueous chromium-containing solution (Palene 60, trademark of Nihon Parkerizing Company, Limited) for 60 seconds at room temperature and then dried in an oven for 20 minutes at 180°C. The sample was subsequently submitted to the various tests, and the results of the performance testing of the test sheet are reported in Table 2.

Test Methods and Evaluation Standards

1. Film Adherence (checkerboard adhesive tape peeling test)

A grid of one hundred 1 mm x 1 mm squares was cut in the test sheet and peeled with adhesive tape, and the number of remaining film squares was counted. This test was conducted both before and after immersion in water at 40°C for 240 hours. The pre-immersion results are reported in row a and the post-immersion results are reported in row b.

2. Film Adherence (Dupont impact test)

A 1 kg impacter (diameter = 1/2 inch) was dropped on the test sheet and the impact area was peeled with adhesive tape. The value measured was the highest impacter drop-height at which film peeling did not occur. This test was conducted both before and after immersion in water at 40°C for 240 hours. The pre-immersion results are reported in row a and the post-immersion results are reported in row b.

3. Corrosion Resistance

The test sheet was prepared by cutting a cross in the coating to reach the base metal, and the test sheet was then subjected to salt-spray testing in accordance with JIS Z-2371 (500 hours for the cold-rolled steel sheet, galvanized hot-dipped zinc-plated steel sheet, and electrogalvanized steel sheet and 1,000 hours for the aluminum sheet). The test sheet was subsequently subjected to tape peeling, and the peel width from the cross cut (one side, mm) was measured.

The aqueous autodeposition coating composition of the present invention is environmentally advantageous because it does not contain chromium. Moreover, coatings with a much better adherence and corrosion resistance than the coatings obtained from prior-art coating compositions are formed by contacting metal surfaces with the aqueous autodeposition coating composition of the present invention. Furthermore, this is achieved without having to carry out a rinse treatment, for example, with a

chromium-containing solution, prior to film curing.

TABLE I

Components In The Aqueous Coating Composition (g/L)				
Production Example No.	A	B	C	D
RESIN (solids)	50	50	100	80
Zn ²⁺ (as zinc)				
Co ²⁺ (as cobalt)				
Mn ²⁺ (as manganese)		0.50		
Ni ²⁺ (as nickel)		1.00		
Fe ²⁺ (as iron)			0.90	
Fe ³⁺ (as iron)	1.50			3.00
Al ³⁺ (as aluminum)				
FLUORIDE OR COMPLEX FLUORIDE				
hydrogen fluoride (as F)	1.20	1.20	1.60	2.70
boron hydrogen fluoride (as F)				0.51
lithium borofluoride (as F)				
zirconium hydrogen fluoride (as F)				
potassium fluozirconate (as F)				
titanium ammonium fluoride (as F)			0.20	
sodium silicofluoride (as F)				
INORGANIC ACIDS				
nitric acid		0.30		
phosphoric acid				
boric acid		0.40	0.60	
ORGANIC ACIDS				
phytic acid				
tannic acid				
TUNGSTEN AND MOLYBDENUM COMPOUNDS				
tungstic acid (as WO ₆)				
Na tungstate (as WO ₆)	0.65	0.60	0.60	0.80
Ca tungstate (as WO ₆)				
K tungstate (as WO ₆)				
molybdic acid (as MoO ₆)				
Na molybdate (as MoO ₆)				
Ca molybdate (as MoO ₆)				
K molybdate (as MoO ₆)				
ADDITIVES				
oxidant (hydrogen peroxide)				
black pigment (solids)	1.40	1.40		
pH	3.2	3.3	3.4	3.6

TABLE I

Components In The Aqueous Coating Composition (g/L)				
Production Example No.	E	F	G	H
RESIN (solids)	50	50	200	50
Zn ²⁺ (as zinc)		0.50		0.60
Co ²⁺ (as cobalt)				
Mn ²⁺ (as manganese)			1.00	
Ni ²⁺ (as nickel)			0.45	0.45
Fe ²⁺ (as iron)	0.40			
Fe ³⁺ (as iron)				
Al ³⁺ (as aluminum)	0.45			
FLUORIDE OR COMPLEX FLUORIDE				
hydrogen fluoride (as F)	2.50	0.50	1.00	1.00
boron hydrogen fluoride (as F)				
lithium borofluoride (as F)	0.16			
zirconium hydrogen fluoride (as F)		0.60		
potassium fluozirconate (as F)		0.25		
titanium ammonium fluoride (as F)				
sodium silicofluoride (as F)			0.40	
INORGANIC ACIDS				
nitric acid				0.40
phosphoric acid				
boric acid	0.80			0.20
ORGANIC ACIDS				
phytic acid				
tannic acid				
TUNGSTEN AND MOLYBDENUM COMPOUNDS				
tungstic acid (as WO ₆)				0.32
Na tungstate (as WO ₆)		0.60		
Ca tungstate (as WO ₆)			0.66	
K tungstate (as WO ₆)				
molybdic acid (as MoO ₆)				0.33
Na molybdate (as MoO ₆)	0.54			
Ca molybdate (as MoO ₆)				
K molybdate (as MoO ₆)				
ADDITIVES				
oxidant (hydrogen peroxide)				
black pigment (solids)				
pH	3.5	2.9	3.3	3.4

TABLE I

Components In The Aqueous Coating Composition (g/L)				
Production Example No.	I	J	K	L
RESIN (solids)	50	50	50	50
Zn ²⁺ (as zinc)				0.80
Co ²⁺ (as cobalt)	0.60		0.45	0.45
Mn ²⁺ (as magnesium)				
Ni ²⁺ (as nickel)		0.70	0.80	
Fe ²⁺ (as iron)				
Fe ³⁺ (as iron)	0.30	0.30	0.25	0.25
Al ³⁺ (as aluminum)				
FLUORIDE OR COMPLEX FLUORIDE				
hydrogen fluoride (as F)	1.00	1.00	1.00	1.00
boron hydrogen fluoride (as F)				
lithium borofluoride (as F)				
zirconium hydrogen fluoride (as F)				
potassium fluozirconate (as F)				
titanium ammonium fluoride (as F)				
sodium silicofluoride (as F)				
INORGANIC ACIDS				
nitric acid	0.50	0.50	0.40	0.40
phosphoric acid				
boric acid	0.20	0.20	0.20	0.20
ORGANIC ACIDS				
phytic acid				
tannic acid				
TUNGSTEN AND MOLYBDENUM COMPOUNDS				
tungstic acid (as WO ₄)				
Na tungstate (as WO ₄)	0.32	0.32		
Ca tungstate (as WO ₄)				0.33
K tungstate (as WO ₄)	0.10			0.20
molybdic acid (as MoO ₄)				
Na molybdate (as MoO ₄)				
Ca molybdate (as MoO ₄)			0.40	
K molybdate (as MoO ₄)		0.10	0.31	
ADDITIVES				
oxidant (hydrogen peroxide)				0.30
black pigment (solids)				
pH	3.5	3.6	3.4	3.3

TABLE I

Components In The Aqueous Coating Composition (g/L)				
Production Example No.	M	N	O	P
RESIN (solids)	80	50	100	100
Zn ²⁺ (as zinc)	0.60		0.50	
Co ²⁺ (as cobalt)				
Mn ²⁺ (as manganese)		0.25		
Ni ²⁺ (as nickel)		1.00		
Fe ²⁺ (as iron)				
Fe ³⁺ (as iron)				0.50
Al ³⁺ (as aluminum)				
FLUORIDE OR COMPLEX FLUORIDE				
hydrogen fluoride (as F)	1.20	0.65	1.50	1.80
boron hydrogen fluoride (as F)				
lithium borofluoride (as F)				
zirconium hydrogen fluoride (as F)				
potassium fluozirconate (as F)				
titanium ammonium fluoride (as F)				
sodium silicofluoride (as F)				
INORGANIC ACIDS				
nitric acid	0.30		0.35	0.35
phosphoric acid	1.45	1.00		
boric acid			0.40	0.40
ORGANIC ACIDS				
phytic acid			0.24	
tannic acid				0.40
TUNGSTEN AND MOLYBDENUM COMPOUNDS				
tungstic acid (as WO ₆)				
Na tungstate (as WO ₆)	0.39	0.60	0.60	0.60
Ca tungstate (as WO ₆)				
K tungstate (as WO ₆)				
molybdic acid (as MoO ₆)				
Na molybdate (as MoO ₆)				
Ca molybdate (as MoO ₆)				
K molybdate (as MoO ₆)				
ADDITIVES				
oxidant (hydrogen peroxide)	0.50			
black pigment (solids)				0.50
pH	3.0	3.1	3.3	3.5

TABLE I

Components In The Aqueous Coating Composition (g/L)	
Production Example No.	Q
RESIN (solids)	50
Zn ²⁺ (as zinc)	
Co ²⁺ (as cobalt)	
Mn ²⁺ (as manganese)	
Ni ²⁺ (as nickel)	
Fe ²⁺ (as iron)	
Fe ³⁺ (as iron)	1.50
Al ³⁺ (as aluminum)	
FLUORIDE OR COMPLEX FLUORIDE	
hydrogen fluoride (as F)	1.20
boron hydrogen fluoride (as F)	
lithium borofluoride (as F)	
zirconium hydrogen fluoride (as F)	
potassium fluozirconate (as F)	
titanium ammonium fluoride (as F)	
sodium silicofluoride (as F)	
INORGANIC ACIDS	
nitric acid	
phosphoric acid	
boric acid	
ORGANIC ACIDS	
phytic acid	
tannic acid	
TUNGSTEN AND MOLYBDENUM COMPOUNDS	
tungstic acid (as WO ₄)	
Na tungstate (as WO ₄)	
Ca tungstate (as WO ₄)	
K tungstate (as WO ₄)	
molybdic acid (as MoO ₄)	
Na molybdate (as MoO ₄)	
Ca molybdate (as MoO ₄)	
K molybdate (as MoO ₄)	
ADDITIVES	
oxidant (hydrogen peroxide)	
black pigment (solids)	1.40
pH	3.3

TABLE 2

	Example					
	1	2	3	4	5	6
Composition Production Number	A	B	C	D	E	F
COATING SUBSTRATE						
cold-rolled steel sheet	+		+	+		+
galvannealed hot-dipped zinc-plated steel sheet						
electrogalvanized steel sheet						
aluminum sheet		+			+	
Chromium Rinse Treatment						
Coating Add-on (g/m ²)	20	15	20	18	13	15
Metal Add-on (mg/m ²)*	205	156	195	220	200	162
Cr Add-on (mg/m ²)						
COATING PROPERTIES						
Adherence (checkerboard peeling)						
a	100	100	100	100	100	100
b	100	100	100	100	100	100
Adherence (Dupont impact)						
a	40	40	40	40	40	40
b	30	40	30	40	40	30
Corrosion Resistance	1.0	0.5	1.5	1.5	1.0	2.0

*total add-on of metal(s) present in the composition

TABLE 2

	Example					
	7	8	9	10	11	12
Composition Production Number	G	H	I	J	K	L
COATING SUBSTRATE cold-rolled steel sheet						
galvannealed hot-dipped zinc-plated steel sheet			+			+
electrogalvanized steel sheet				+		
aluminum sheet	+	+				+
Chromium Rinse Treatment						
Coating Add-on (g/m ²)	17	15	14	15	13	12
Metal Add-on (mg/m ²)*	138	125	155	140	135	120
Cr Add-on (mg/m ²)						
COATING PROPERTIES Adherence (checkerboard peeling) a	100	100	100	100	100	100
b	100	100	100	100	100	100
Adherence (Dupont impact) a	40	40	40	40	30	40
b	40	30	30	40	30	30
Corrosion Resistance	0	0	1.5	1.5	1.5	0.5

*total add-on of metal(s) present in the composition

TABLE 2

	Example			
	13	14	15	16
Composition Production Number	M	N	O	P
COATING SUBSTRATE				
cold-rolled steel sheet	+			
galvannealed hot-dipped zinc-plated steel sheet				
electrogalvanized steel sheet				+
aluminum sheet		+	+	
Chromium Rinse Treatment				
Coating Add-on (g/m ²)	16	12	14	20
Metal Add-on (mg/m ²)	180	140	160	205
Cr Add-on (mg/m ²)				
COATING PROPERTIES				
Adherence (checkerboard peeling)				
a	100	100	100	100
b	100	100	100	100
Adherence (Dupont impact)				
a	40	40	40	40
b	30	30	40	40
Corrosion Resistance	1.5	0.5	0	1.0

*total add-on of metal(s) present in the composition

TABLE 2

	Comparison Example			
	1	2	3	4
Composition Production Number	Q			
COATING SUBSTRATE cold-rolled steel sheet	+			+
galvannealed hot-dipped zinc-plated steel sheet		+		
electrogalvanized steel sheet			+	
aluminum sheet				
Chromium Rinse Treatment				+
Coating Add-on (g/m^2)	24	20	20	25
Metal Add-on (mg/m^2)	235	195	210	205
Cr Add-on (mg/m^2)				105
COATING PROPERTIES Adherence (checkerboard peeling)				
a	100	100	100	100
b	100	100	100	100
Adherence (Dupont impact)				
a	20	30	30	20
b	20	20	20	20
Corrosion Resistance	3.5	4.0	4.0	3.0

*total add-on of metal(s) present in the composition

We claim

1. An aqueous autodeposition coating composition having a pH in the range of about 1.6 to about 5.0 which comprises:
 - 5 a) a water dispersible or water-soluble organic film forming resin;
 - b) at least one of fluoride ion or fluoride ion and complex fluoride ion;
 - c) additional metal ions of at least one metal
10 selected from the group consisting of zinc, cobalt, manganese, nickel, iron and aluminum;
 - d) at least one of tungstate ion and molybdate ion; and
 - e) water
- 15 2. An aqueous autodeposition coating composition of claim 1 containing a pH adjusting amount of at least one acid selected from the group consisting of nitric and, phosphoric acid, boric acid, phytic acid and tannic acid.
- 20 3. An autodeposition coating composition of claim 1 wherein the organic film forming resin comprises at least one film forming resin selected from the group consisting of urethane resin, epoxy resin, polyester resin, and a resin comprising a residue of at least
25 one monomer selected from the group consisting of methyl acrylate, ethyl acrylate or butyl acrylate, 2-hydroxy ethyl acrylate, 2-hydroxypropyl acrylate, 2-ethylhexyl acrylate, methyl methacrylate, ethyl

methacrylate, n-butyl methacrylate, 2-hydroxyethyl methacrylate, 2-hydroxypropyl methacrylate, glycidyl acrylate, glycidyl methacrylate, acrylamide, methacrylamide, ethylene, styrene, vinyl chloride, vinylidene chloride, vinyl acetate, acrylic acid and methacrylic acid.

4. An autodeposition coating composition of claim 1 containing from 5 to 550 grams/liter of film forming resin.
- 10 5. An autodeposition coating composition of claim 2 containing from 5 to 550 grams/liter of film forming resin.
6. An autodeposition coating composition of claim 1 containing from 0.1 to 5 grams per liter of fluoride ion or fluoride ion and complex fluoride ion measured as fluorine.
- 15 7. An autodeposition coating composition of claim 1 containing 0.1 to 20 grams/liter of additional metal ions.
- 20 8. An autodeposition coating composition of claim 1 containing from 0.1 to 5 grams/liter of at least one of tungstate ion and molybdate ion.
9. An autodeposition coating composition of claim 1 which comprises:
 - 25 a) 5 to 550 grams/liter of the water soluble or water dispersible film forming resin;
 - b) 0.1 to 5 grams/liter of fluoride ion or fluoride ion and complex fluoride ion measured as

fluorine;

- c) 0.1 to 20 grams/liter additional metal ions; and
- d) 0.1 to 5 grams/liter of at least one of tungstate ion and molybdate ion.

5 10. An autodeposition coating of claim 2 comprising:

- a) 5 to 550 grams/liter of the water soluble or water dispersible film forming resin;
- b) 0.1 to 5 grams/liter of fluoride ion or fluoride ion and complex fluoride ion measured as
10 fluorine;

- c) 0.1 to 20 grams/liter additional metal ions;
- d) 0.1 to 5 grams/liter of at least one of tungstate ion and molybdate ion.

11. An autodeposition coating of claim 3 which comprises:

- a) 5 to 550 grams/liter of the water soluble or water dispersible film forming resin;
- b) 0.1 to 5 grams/liter of fluoride ion or fluoride ion and complex fluoride ion measured as
15 fluorine;

- c) 0.1 to 20 grams/liter additional metal ions;
- d) 0.1 to 5 grams/liter of at least one of tungstate ion and molybdate ion.

12. An autodeposition coating of claim 9 wherein the organic film forming resin is present at from 50 to
25 100 grams/liter.

13. An autodeposition coating of claim 9 containing from 0.5 to 3 grams/liter of fluoride ion or fluoride ion and complex fluoride ion measured as fluorine.

14. A composition of claim 9 containing from 0.3 to 6 grams per liter of additional metal ions.
15. A composition of claim 11 containing from 0.5 to 3 grams/liter of fluoride ions or fluoride ion and complex fluoride ions measured as fluorine.
- 5 16. A composition of claim 9 which comprises
- a) 50 to 100 grams/liter organic film forming resin;
 - b) 0.5 to 3 grams/liter fluoride ion or fluoride ion and complex fluoride ion measured as fluorine;
 - 10 c) 0.3 to 6 grams/liter additional metal ions; and
 - d) 0.3 to 2 grams/liter of at least one of tungstate.
17. A composition of claim 16 wherein the organic film forming resin comprises at least one film forming resin selected from the group consisting of urethane resin, epoxy resin, polyester resin, and a resin containing a residue of at least one monomer selected from the group consisting of methyl acrylate, ethyl acrylate or butyl acrylate, 2-hydroxy ethyl acrylate, 2-hydroxypropyl acrylate, 2-ethylhexyl acrylate, methyl methacrylate, ethyl methacrylate, n-butyl methacrylate, 2-hydroxyethyl methacrylate, 2-hydroxypropyl methacrylate, glycidyl acrylate, glycidyl methacrylate, acrylamide, methacrylamide, ethylene, styrene, vinyl chloride, vinylidene chloride, vinyl acetate, acrylic acid and methacrylic acid.
- 15 20 25

INTERNATIONAL SEARCH REPORT

PCT/US 93/00137

International Application No

I. CLASSIFICATION OF SUBJECT MATTER (if several classification symbols apply, indicate all) ⁶		
According to International Patent Classification (IPC) or to both National Classification and IPC Int.Cl. 5 C09D5/08		
II. FIELDS SEARCHED		
Minimum Documentation Searched ⁷		
Classification System	Classification Symbols	
Int.Cl. 5	C09D	
Documentation Searched other than Minimum Documentation to the Extent that such Documents are Included in the Fields Searched ⁸		
III. DOCUMENTS CONSIDERED TO BE RELEVANT⁹		
Category ¹⁰	Citation of Document, ¹¹ with indication, where appropriate, of the relevant passages ¹²	Relevant to Claim No. ¹³
X	EP,A,0 046 268 (UNION CARBIDE CORPORATION) 24 February 1982 see page 6, line 7 - page 8, line 21 see example 2 see page 22, line 5 - line 6 see page 22, line 17 ---	1-17
A	EP,A,0 312 648 (AMCHEM PRODUCTS, INC.) 26 April 1989 see abstract see example 1 ---	1-17
A	GB,A,1 467 151 (AMCHEM PRODUCTS) 16 March 1977 cited in the application see claims 1,7-15 ---	1-17
		-/--
¹⁰ Special categories of cited documents : "A" document defining the general state of the art which is not considered to be of particular relevance "E" earlier document but published on or after the international filing date "L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified) "O" document referring to an oral disclosure, use, exhibition or other means "P" document published prior to the international filing date but later than the priority date claimed "T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention "X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art "&" document member of the same patent family		
IV. CERTIFICATION		
Date of the Actual Completion of the International Search 16 MARCH 1993		Date of Mailing of this International Search Report 07.04.93
International Searching Authority EUROPEAN PATENT OFFICE		Signature of Authorized Officer GOERKE H.R.

III. DOCUMENTS CONSIDERED TO BE RELEVANT (CONTINUED FROM THE SECOND SHEET)		
Category °	Citation of Document, with indication, where appropriate, of the relevant passages	Relevant to Claim No.
A	EP,A,0 005 158 (VIANOVA KUNSTHARZ AG) 14 November 1979 see abstract see page 3, line 20 - line 24 -----	1-17

**ANNEX TO THE INTERNATIONAL SEARCH REPORT
ON INTERNATIONAL PATENT APPLICATION NO.**

US 9300137
SA 68899

This annex lists the patent family members relating to the patent documents cited in the above-mentioned international search report.
The members are as contained in the European Patent Office EDP file on
The European Patent Office is in no way liable for these particulars which are merely given for the purpose of information. 16/03/93

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
EP-A-0046268	24-02-82	JP-A- 57059965	10-04-82
		JP-B- 61047181	17-10-86
EP-A-0312648	26-04-89	AU-B- 575333	28-07-88
		AU-A- 3112584	31-01-85
		CA-A- 1280843	26-02-91
		EP-A, B 0132828	13-02-85
		EP-A- 0310708	12-04-89
		EP-A- 0310709	12-04-89
		JP-A- 60058474	04-04-85
		US-A- 4562098	31-12-85
		US-A- 4647480	03-03-87
GB-A-1467151	16-03-77	JP-C- 963678	20-07-79
		JP-A- 49113831	30-10-74
		JP-B- 53044949	02-12-78
		AU-B- 476261	16-09-76
		AU-A- 6619174	04-09-75
		BE-A- 811841	01-07-74
		CA-A- 1043922	05-12-78
		SE-B- 415823	03-11-80
		US-A- 4103049	25-07-78
		US-A- 4160756	10-07-79
EP-A-0005158	14-11-79	AT-A- 352843	10-10-79
		CA-A- 1144673	12-04-83
		JP-A- 54143444	08-11-79
		US-A- 4239672	16-12-80