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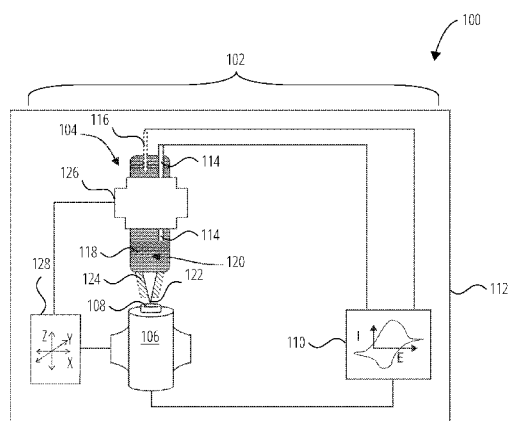


FIG. 1

(57) Abstract: An electrochemical system includes at least one electrochemical cell with a receptacle containing an electrolytic bath in which is disposed a counter electrode. At least one nozzle opens from the receptacle toward and proximate a substrate configured as a working electrode. The at least one electrochemical cell is selectively configurable between a configuration for electrodeposition of a material onto the substrate and a configuration for electrodis-solution of material from a structure on the substrate. In a method of using an electrochemical cell, a metal salt--of the electrolytic bath--is flowed through the nozzle in the presence of at least one of a voltage difference and a current flow between the working electrode and the counter electrode. The system may be configured for relative movement between the at least one nozzle and the substrate, and the electrochemical cell(s) may be usable for any of elec-trodeposition, electrodis-solution, and electrochemical sensing.

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TITLE

THREE-DIMENSIONAL ELECTROCHEMICAL MANUFACTURING AND SENSING
SYSTEM AND RELATED METHODS

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PRIORITY CLAIM

This application claims the benefit of the filing date of United States Provisional
Patent Application Serial No. 62/777,843, filed December 11, 2018, for “Three-Dimensional
10 Electrochemical Manufacturing and Sensing System and Related Methods.”

STATEMENT REGARDING FEDERALLY SPONSORED RESEARCH OR DEVELOPMENT

This invention was made with government support under Contract Number
15 DE-AC07-05-ID14517 awarded by the United States Department of Energy. The
government has certain rights in the invention.

TECHNICAL FIELD

Embodiments of the disclosure relate generally to systems and methods for
20 performing electrochemical reactions and processes. More particularly, embodiments of the
disclosure relate to systems for performing electrodeposition, electrodisolution, and
electrochemical sensing of three-dimensional structures.

BACKGROUND

25 Nuclear reactors are used to generate power (*e.g.*, electrical power) using nuclear fuel
materials. For example, heat generated by nuclear reactions carried out within the nuclear
fuel materials may be used to boil water, and the steam resulting from the boiling water may
be used to rotate a turbine. Rotation of the turbine may be used to operate a generator for
generating electrical power.

30 Nuclear reactors generally include what is referred to as a “nuclear core,” which is the
portion of the nuclear reactor that includes the nuclear fuel material and is used to generate
heat from the nuclear reactions of the nuclear fuel material. The nuclear core may include a
plurality of fuel rods, which include the nuclear fuel material.

Most nuclear fuel materials include one or more of the elements of uranium and
35 plutonium (although other elements such as thorium are also being investigated). There are,

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however, different types or forms of nuclear fuel materials that include such elements. For example, nuclear fuel pellets may comprise ceramic nuclear fuel materials. Ceramic nuclear fuel materials include, among others, radioactive uranium oxide (*e.g.*, uranium dioxide (UO₂), which is often abbreviated as “UOX”), which is often used to form nuclear fuel pellets.

- 5 Mixed oxide radioactive ceramic materials (which are often abbreviated as “MOX”) are also commonly used to form nuclear fuel pellets. Such mixed oxide radioactive ceramic materials may include, for example, a blend of plutonium oxide and uranium oxide. Such a mixed oxide may include, for example, U_{1-x}Pu_xO₂, wherein *x* is between about 0.2 and about 0.3. Transuranic (TRU) mixed oxide radioactive ceramic materials (which are often
- 10 abbreviated as “TRU-MOX”) also may be used to form nuclear fuel pellets. Transuranic mixed oxide radioactive ceramic materials include relatively higher concentrations of minor actinides such as, for example, neptunium (Np), americium (Am), and curium (Cm). Carbide nuclear fuels and mixed carbide nuclear fuels having compositions similar to the oxides mentioned above, but wherein carbon is substituted for oxygen, are also being investigated
- 15 for use in nuclear reactors.

In addition to ceramic nuclear fuel materials, there are also metallic nuclear fuel materials. Metallic nuclear fuels include, for example, metals based on one or more of uranium, plutonium, and thorium. Other elements such as hydrogen (H), zirconium (Zr), molybdenum (Mo), and others may be incorporated in uranium- and plutonium-based metals.

- 20 In nuclear reactors that employ metallic nuclear fuels, the metallic nuclear fuel is often formed into rods or pellets of predetermined size and shape (*e.g.*, spherical, cubical, cylindrical, *etc.*) that at least substantially comprise the metallic nuclear fuel. The nuclear fuel material is contained within and at least partially surrounded by a cladding material, which may be in the form of, for example, an elongated tube. The cladding material is used
- 25 to hold and contain the nuclear fuel. The cladding material typically comprises a metal or metallic alloy, such as stainless steel. During operation of the nuclear reactor, the cladding material may separate (*e.g.*, isolate and hermetically seal) the nuclear fuel bodies from a liquid (*e.g.*, water or molten salt) that is used to absorb and transport the heat generated by the nuclear reaction occurring within the nuclear fuel.

- 30 Traditional methods of manufacturing the foregoing nuclear fuel materials include the processing of nuclear fuel powders using so-called dry or wet processes and/or using high temperature (*e.g.*, 1600°C or greater) melting or laser-beam melting. Such traditional methods result in significant safety and environmental concerns. For example, such high

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temperature and laser-beam melting processes are associated with high energy expenditures. The dispersion of radioactive nuclear fuel powders to the atmosphere during manufacturing of the nuclear fuel materials also poses a significant safety risk. Traditional machining processes may also include one or more machining steps or leaching steps to remove material from the nuclear fuel materials, and the machining and/or leaching steps generate material waste. Thus, improved systems and methods of manufacturing nuclear fuels that reduce costs, waste, and safety risks are desirable.

DISCLOSURE

An electrochemical system, according to embodiments of the disclosure, comprises at least one electrochemical cell. The at least one electrochemical cell comprises a receptacle containing an electrolytic bath. At least one nozzle opens from the receptacle toward and proximate a substrate configured as a working electrode of the at least one electrochemical cell. A counter electrode is disposed in the electrolytic bath. The at least one electrochemical cell is selectively configurable between a configuration for electrodeposition of material to form a structure on the substrate and a configuration for electrodisolution of material from the structure on the substrate.

A method of using an electrochemical cell, according to embodiments of the disclosure, comprises providing an electrolytic bath in a receptacle. The electrolytic bath comprises a metal salt. A counter electrode is disposed at least partially within the electrolytic bath. The counter electrode is coupled to a working electrode. In the presence of at least one of a voltage difference and a current flow between the working electrode and the counter electrode, the metal salt is flowed through a nozzle coupled to the receptacle.

BRIEF DESCRIPTION OF THE DRAWINGS

FIG. 1 is a schematic representation of an electrochemical system, according to embodiments of the disclosure, wherein the system includes an electrochemical cell and at least one controller.

FIG. 2 is a schematic representation of an electrochemical system, according to embodiments of the disclosure, wherein the system includes an electrochemical cell and at least two controllers.

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FIG. 3 is a schematic representation of an electrochemical system, according to embodiments of the disclosure, wherein the system includes an electrochemical cell and at least three controllers.

FIG. 4 is a schematic representation of a system with a plurality of electrochemical cells, according to embodiments of the disclosure, which plurality of electrochemical cells may be incorporated within an electrochemical system, such as the electrochemical system of any one FIG. 1, FIG. 2, and/or FIG. 3.

FIG. 5 is a schematic, cross-sectional, elevational representation of a nuclear fuel element formed by electrodeposition using the system of any of FIG. 1, FIG. 2, FIG. 3, and/or FIG. 4, according to embodiments of the disclosure.

FIG. 6A and FIG. 6B are schematic polarization curves for the electrodeposition of compounds using the system of any of FIG. 1, FIG. 2, FIG. 3, and/or FIG. 4.

FIG. 7 is a schematic representation of an electrochemical cell configured as an electrodeposition cell, which may be incorporated in the system of any of FIG. 1, FIG. 2, FIG. 3, and/or FIG. 4, according to embodiments of the disclosure.

FIG. 8 is a schematic representation of an electrochemical cell configured as an electrodisolution cell, which may be incorporated in the system of any of FIG. 1, FIG. 2, FIG. 3, and/or FIG. 4, according to embodiments of the disclosure.

FIG. 9 is a schematic representation of an electrochemical cell configured as an electrochemical-sensor cell, which may be incorporated in the system of any of FIG. 1, FIG. 2, FIG. 3, and/or FIG. 4, according to embodiments of the disclosure.\

MODE(S) FOR CARRYING OUT THE INVENTION

Systems and methods disclosed herein enable fabrication of three-dimensional structures, such as nuclear fuel elements, by additive manufacturing through electrodeposition using at least one electrochemical cell; etching of material from a structure through electrodisolution using at least one electrochemical cell; and/or sensing of a characteristic or composition of a material through electrochemical sensing using at least one electrochemical cell. The electrodeposition, electrodisolution, and/or electrochemical sensing of, *e.g.*, nuclear material, may be accomplished at relatively low temperatures, with less risk of dispersion of radioactive nuclear fuel material into the atmosphere, with less material waste, with less energy expenditure, with less expense, and with increased safety.

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The following description provides specific details, such as compositions, materials, processing conditions, equipment, and features thereof, in order to provide a thorough description of embodiments described herein. However, a person of ordinary skill in the art will understand that the embodiments disclosed herein may be practiced without employing these specific details. Indeed, the embodiments of the disclosure may be practiced in conjunction with conventional fabrication techniques employed in the industry. In addition, the description provided below does not form a complete process flow, apparatus, or system for forming a component of a nuclear reactor, another structure, or related methods. Only those process acts and structures necessary to understand the embodiments of the disclosure are described in detail below. Additional acts to form a component of a nuclear reactor core or another structure may be performed by conventional techniques. Further, any drawings accompanying the present application are for illustrative purposes only and, thus, are not necessarily drawn to scale.

The illustrations included herewith are not meant to be actual views of any particular systems or structures formed with the systems, but are merely idealized representations that are employed to describe embodiments herein. Elements and features common between figures may retain the same numerical designation.

As used herein, the terms “comprising,” “including,” “containing,” “characterized by,” and grammatical equivalents thereof are inclusive or open-ended terms that do not exclude additional, unrecited elements or method steps, but also include the more restrictive terms “consisting of” and “consisting essentially of” and grammatical equivalents thereof.

As used herein, the term “may” with respect to a material, structure, feature, or method act indicates that such is contemplated for use in implementation of an embodiment of the disclosure, and such term is used in preference to the more restrictive term “is” so as to avoid any implication that other compatible materials, structures, features and methods usable in combination therewith should or must be excluded.

As used herein, the term “configured” refers to a size, shape, material composition, and arrangement of one or more of at least one structure and at least one apparatus facilitating operation of one or more of the structure and the apparatus in a predetermined way.

As used herein, the term “substantially” in reference to a given parameter, property, or condition means and includes to a degree that one of ordinary skill in the art would understand that the given parameter, property, or condition is met with a degree of variance, such as within acceptable manufacturing tolerances. By way of example, depending on the

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particular parameter, property, or condition that is substantially met, the parameter, property, or condition may be at least 90.0% met, at least 95.0% met, at least 99.0% met, even at least 99.9% met, or even 100.0% met.

As used herein, “about” or “approximately” in reference to a numerical value for a particular parameter is inclusive of the numerical value and a degree of variance from the numerical value that one of ordinary skill in the art would understand is within acceptable tolerances for the particular parameter. For example, “about” or “approximately” in reference to a numerical value may include additional numerical values within a range of from 90.0% to 110.0% of the numerical value, such as within a range of from 95.0% to 105.0% of the numerical value, within a range of from 97.5% to 102.5% of the numerical value, within a range of from 99.0% to 101.0% of the numerical value, within a range of from 99.5% to 100.5% of the numerical value, or within a range of from 99.9% to 100.1% of the numerical value.

As used herein, “and/or” includes any and all combinations of one or more of the associated listed items.

Embodiments of the disclosure relate to systems and related methods for manufacturing (*e.g.*, depositing, forming), machining (*e.g.*, shaping, finishing), and analyzing composition of a structure (*e.g.*, a three-dimensional structure). FIG. 1 illustrates a schematic of a system 100, according to embodiments of the disclosure. The system 100 comprises an electrochemical processing unit 102, which includes an electrochemical cell 104 that includes a substrate 106 (*e.g.*, a platform) on which a three-dimensional (3D) structure 108 may be formed. The electrochemical processing unit 102 also comprises at least one controller (*e.g.*, controller 110), such as a computer. One or more of the components of the electrochemical processing unit 102, such as one or more of the electrochemical cell 104, the substrate 106 thereof, the structure 108, and the controller 110 may be enclosed within a reaction chamber 112 (*e.g.*, a radioactive shield).

The electrochemical cell 104 of the electrochemical processing unit 102 includes multiple electrodes. The substrate 106 of the electrochemical cell 104 serves as a working electrode. A counter electrode 114 is also included and, in some embodiments, also a reference electrode 116.

The electrochemical cell 104 of the electrochemical processing unit 102 further includes a container 118 (*e.g.*, a receptacle), such as a crucible, in which an electrolytic bath 120 is retained. The reference electrode 116, if included, and the counter electrode 114 may

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be at least partially disposed in the electrolytic bath 120. At least one nozzle 122 may be coupled to the container 118. In some embodiments, a heater 124 (*e.g.*, an induction heater or a heating block, either of which can be controlled by a temperature control unit) may be coupled to and disposed about the nozzle 122 and/or about the substrate 106 (*e.g.*, the working electrode). In some embodiments, the heater 124 may comprise an induction heater that laterally surrounds each nozzle 122.

The substrate 106 (*e.g.*, the working electrode) may be disposed proximate to the nozzle 122 such that one or more elements of the electrolytic bath 120 may be passed through the nozzle 122 (*e.g.*, deposited from the nozzle 122 and onto a surface of the substrate 106; dissolved up into the nozzle 122 from the structure 108 on the substrate 106). Another container (not illustrated) may be included in the electrochemical processing unit 102 and may be proximate at least the surface of the substrate 106; the structure 108 during its formation, dissolution, or analyzing; and at least a lowest part of the nozzle 122. Such other container may be formed of steel, glass, plastic, or the like.

One or more of the substrate 106, the counter electrode 114, the reference electrode 116 (if included), and the nozzle 122 may be selected to comprise silver, titanium, gold, and/or a boron-containing material such as borosilicate glass, boron carbide, and high-boron steel. In some embodiments, the counter electrode 114 may be selected to comprise a metal substantially similar to a composition of a metal to be deposited to (or removed by dissolution from) the structure 108, using the system 100, as described further, below, with reference to FIG. 4, FIG. 7, and FIG. 8. In further embodiments, one or more of the substrate 106 (*e.g.*, the working electrode), the counter electrode 114, and the reference electrode 116 (if included) may be selected to comprise a material compatible with a composition (*e.g.*, chemistry) of the structure 108 being fabricated (*e.g.*, by electrodeposition), machined (*e.g.*, etched through electrodisolution), or analyzed (*e.g.*, through electrochemical sensing) using the system 100.

In a method for using the system 100, according to embodiments of the disclosure, a voltage differential is selected and is selectively applied by the controller 110 and an electrical current is selected and selectively flowed between the substrate 106 (*e.g.*, the working electrode) and the counter electrode 114 such that the system 100 may be used for electrochemical manufacturing in a fashion that allows for material deposition (*e.g.*, by electrodeposition), material dissolution (*e.g.*, by electrodisolution), and/or analysis (*e.g.*, by electrochemical sensing of the electrochemical state of the system) with additional

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electromechanical control of the position of the substrate 106, the counter electrode 114, and the nozzle 122 of the container 118. In some embodiments, a voltage differential is selected and applied by the controller 110 such that a proportional (*e.g.*, corresponding) current flows between the substrate 106 (*e.g.*, the working electrode) and the counter electrode 114. In
5 other embodiments, a current is selected and is flowed by the controller 110 and a proportional voltage differential is applied between the substrate 106 (*e.g.*, the working electrode) and the counter electrode 114.

With regard to the electromechanical control of the position of the substrate 106, the counter electrode 114, and the nozzle 122 of the container 118, one or both of the substrate
10 106 and the container 118 of the electrochemical cell 104 may be coupled to an electromechanical arm 126 such that the substrate 106 and the container 118 may be configured to move in the x-direction (*i.e.*, left and right, along arrow X, in the view illustrated in FIG. 1), the y-direction (*i.e.*, into and out of the page in the view illustrated in FIG. 1), and the z-direction (*i.e.*, up and down, along arrow Z, in the view illustrated in FIG.
15 1). As the container 118 is moved in this fashion, the nozzle 122 is also moved in the same direction, *e.g.*, over the upper surface of the substrate 106 and along the structure 108 supported by the substrate 106.

In some embodiments, the electromechanical arm 126 may also be configured to control movement of the substrate 106 (and therefore also the structure 108), such as by
20 rotating the substrate 106. Accordingly, the electromechanical arm 126 of such embodiments may rotate the substrate 106 and/or the container 118 (and nozzle 122) about any or each axis of movement (*e.g.*, the x-, y-, and z-directions) such that the electromechanical arm 126 may also be able to pitch, roll, etc. The electromechanical arm 126 may be configured to manipulate the movement of the substrate 106 (and therefore also the structure 108) and the
25 container 118 (and therefore also the nozzle 122) either jointly (*e.g.*, as the substrate 106 is moved in a certain direction, the container 118 is also moved in the same direction) or independently (*e.g.*, enabling the substrate 106 to be moved in one direction while the container 118 is motionless or moved in a different direction).

In some embodiments, the electrochemical processing unit 102 of the system 100 also
30 includes an XYZ platform 128 that may support the substrate 106 (and therefore also the structure 108). In such embodiments, the XYZ platform 128 may be configured to be manipulated to control the movement of the substrate 106 (and therefore also the structure

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108), while the electromechanical arm 126 may be dedicated for controlled manipulation of the container 118 (and therefore also the nozzle 122).

At least one of the controllers of the at least one controller of the system 100, *e.g.*, the controller 110 of FIG. 1, may be in operable communication with the electromechanical arm 126. Therefore, the controller 110 may be configured to control the movement of the electromechanical arm 126 and therefore the movement of at least the container 118 and the nozzle 122. In embodiments in which the electromechanical arm 126 is also operatively connected to the substrate 106, the controller 110 may also be configured to control the movement of the substrate 106 and therefore the movement of the structure 108. In other embodiments in which the XYZ platform 128 is included and is operatively connected to the substrate 106, the controller 110 may be configured to control the movement of the XYZ platform 128 and therefore the movement of the substrate 106 and the structure 108.

FIG. 1 illustrates a system (*e.g.*, system 100) with one controller (*e.g.*, controller 110) for controlling the voltage differential and current flow to/from/between the substrate 106 (*e.g.*, the working electrode), the counter electrode 114, and the reference electrode 116 (if included). In other embodiments, however, the more than one controller may be included in the system.

For example, FIG. 2 illustrates a system 200 with an electrochemical processing unit 202 that includes the electrochemical cell 104 and two controllers: a first controller 204 and a second controller 206. The first controller 204 may be configured to control the voltage differential and current flow to/from/between the substrate 106 (*e.g.*, the working electrode), the counter electrode 114, and the reference electrode 116 (if included). The second controller 206 may be configured to control the movement of both the electromechanical arm 126 (and therefore the container 118 and the nozzle 122) and the XYZ platform 128 (and therefore the substrate 106 and the structure 108).

As another example, FIG. 3 illustrates a system 300 with an electrochemical processing unit 302 that includes the electrochemical cell 104 and three controllers: the first controller 204, a second controller 304, and third controller 306. As in the system 200 of FIG. 2, the first controller 204 may be configured to control the voltage differential and current flow to/from/between the substrate 106 (*e.g.*, the working electrode), the counter electrode 114, and the reference electrode 116 (if included). The second controller 304 may be configured to control the movement of the electromechanical arm 126 (and therefore the

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container 118 and the nozzle 122). The third controller 306 may be configured to control the movement of the XYZ platform 128 (and therefore the substrate 106 and the structure 108).

In still other embodiments, one or more additional controllers may be included in the system to control additional system equipment, such as to control the heat applied (*e.g.*, to the nozzle 122) by the heater 124. Alternatively, one or more of the aforementioned controllers (*e.g.*, the controller 110 of the system 100 of FIG. 1; the first controller 204 or the second controller 206 of the system 200 of FIG. 2; or the first controller 204, the second controller 304, or the third controller 306 of the system 300 of FIG. 3) may be additionally configured to control operation of other system equipment, such as the heat applied to the nozzle 122 by the heater 124.

Any or all of the aforementioned controllers (*e.g.*, the controller 110 of the system 100 of FIG. 1; the first controller 204 or the second controller 206 of the system 200 of FIG. 2; or the first controller 204, the second controller 304, or the third controller 306 of the system 300 of FIG. 3) may be or include a potentiostat, a galvanostat, a power source, such as a DC power supply, or other instrumentation to control the operation of the corresponding system component (*e.g.*, with regard to the controller 110 of FIG. 1 or the first controller 204 of FIG. 2 or FIG. 3, to control the current flow and/or a voltage (*e.g.*, potential difference) applied between the substrate 106 (*e.g.*, the working electrode) and the counter electrode 114). As illustrated in FIG. 4, in some embodiments, multiple electrochemical cells 104 may be coupled to any one or more of the controllers of the system.

In some embodiments, the substrate 106 may be supported on (*e.g.*, directly on top of) the XYZ platform 128, as illustrated in FIG. 2 and FIG. 3. In other embodiments, the XYZ platform 128 may be incorporated within (*e.g.*, be integral to) the substrate 106.

In some embodiments, the reaction chamber 112 may comprise a radioactive shield configured to contain radioactive materials that may be used to manufacture (*e.g.*, by electrodeposition), machine (*e.g.*, by electrodisolution), or analyze (*e.g.*, by electrochemical sensing) the structure 108 therein. The reaction chamber 112 may also be configured to provide a controlled environment in which the nuclear fuel element 500 of FIG. 5, described below, may be manufactured.

While the system 100 of FIG. 1, the system 200 of FIG. 2, and the system 300 of FIG. 3 are illustrated as having as having a single electrochemical cell 104, the disclosure is not so limited. As illustrated in FIG. 4, a system 400 may include a plurality of electrochemical cells 104 coupled to one or more controllers, such as the controller 110. The system 400 may

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further comprise the substrate 106 (*e.g.*, the working electrode), which may be a single substrate 106 for use with all of the electrochemical cells 104, as illustrated in FIG. 4. Or, in other embodiments, each electrochemical cell 104 may include a separate substrate 106, or some of the electrochemical cells 104 may share a substrate 106 what others of the electrochemical cells 104 have their own substrate 106. The system 400 may also include one or more electromechanical arm 126 and/or one or more XYZ platform 128 for one, all, or some of the electrochemical cells 104 and/or one, all, or some of the structures 108 being fabricated, machined, or analyzed.

Each of the electrochemical cells 104 may comprise a respective container 118, nozzle 122, and, optionally, a heater 124 (FIG. 1, FIG. 2, FIG. 3). A plurality of the electrochemical cells 104 may be provided within the system 400 (or in instead and in place of the single electrochemical cell 104 of the system 100 of FIG. 1, the system 200 of FIG. 2, or the system 300 of FIG. 3) and within the reaction chamber 112 (FIG. 1, FIG. 2, FIG. 3). In some embodiments, each respective container 118 may contain an electrolytic bath 120 having a different composition (*e.g.*, composition A^{m+} , composition B^{n+} , composition $A^{m+} + B^{n+}$). Accordingly, a plurality of electrochemical cells 104 may concurrently manufacture (*e.g.*, by electrodeposition), machine (*e.g.*, by electrodisolution), or analyze (*e.g.*, by electrochemical sensing) one or more individual structures 108 (*e.g.*, the structure 108 of composition "A," the structure 108 of composition "B," and the structure 108 of composition "AB") or one or more portions (*e.g.*, regions) of the same structure (*e.g.*, concurrently). In other embodiments, more than one nozzle 122 and, optionally, respective heater 124 (FIG. 1, FIG. 2, FIG. 3) may be coupled to a single container 118 of a single electrochemical cell 104.

The composition of the electrolytic bath 120 may be tailored based upon a composition of the structure 108 to be manufactured (*e.g.*, by electrodeposition), to be machined (*e.g.*, by electrodisolution), or to be analyzed (*e.g.*, by electrochemical sensing) using the system (*e.g.*, the system 100 of FIG. 1, the system 200 of FIG. 2, the system 300 of FIG. 3, or the system 400 of FIG. 4). The electrolytic bath 120 of any of the aforementioned electrochemical cells 104 and systems may comprise a room temperature ionic liquid formulated to permit the flow of electricity therein. The ionic liquid may include hydrogen and/or carbon, each of which is capable of providing shielding against gamma and neutron radiation and of preventing the transportation of air-borne radioactive elements, when such radioactive elements are dissolved in the electrolytic bath 120, *e.g.*, for deposition by or as a

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result of dissolution using the system (*e.g.*, system 100 of FIG. 1, system 200 of FIG. 2, system 300 of FIG. 3, system 400 of FIG. 4). In some embodiments, the ionic liquid of the electrolytic bath 120 may comprise nitrogen-containing cations, such as imidazolium and nitrogen-, bromine-, or boron-containing anions, such as dicyanamide anion ($\text{N}(\text{CN})_2^-$),
5 bromine (Br^-), and tetrafluoroborate (BF_4^-). By way of non-limiting example, the electrolytic bath 120 may comprise an imidazolium-based ionic liquid including 1-butyl-3-methylimidazolium tetrafluoroborate or 1-ethyl-3-methylimidazolium bromide. In such embodiments, the ionic liquid composition of the electrolytic bath 120 may have the advantage of higher neutron absorption cross-sections and may offset the moderator effects of
10 hydrogen and/or carbon, which may also be included in the electrolytic bath 120. The electrolytic bath 120 may further comprise an electrolyte, or salt. In some embodiments, the electrolytic bath 120 comprises a metal salt. Such salts may include, for example, AlBr_3 , LiBF_4 , LiBr , KBr , and CsBr . Salts such as LiBr and LiBF_4 may also have the advantage of offsetting the moderator effects of hydrogen and carbon included in the electrolytic bath 120.

15 In some embodiments, the system (*e.g.*, the system 100 of FIG. 1, the system 200 of FIG. 2, the system 300 of FIG. 3, the system 400 of FIG. 4) is operated so as to form, by electrodeposition, a nuclear fuel element such as the nuclear fuel element 500 of FIG. 5. Accordingly the electrolytic bath 120 may have one or more elements, dissolved in the ionic liquid of the electrolytic bath 120, of nuclear material to be included in the nuclear fuel
20 element 500 to be fabricated.

With reference to FIG. 5, illustrated is a nuclear fuel element 500 that may be fabricated in whole or in part using a system (*e.g.*, the system 100 of FIG. 1, the system 200 of FIG. 2, the system 300 of FIG. 3, the system 400 of FIG. 4) and a method of embodiments of the disclosure. FIG. 5 illustrates the nuclear fuel element 500 in elevational cross-section.
25 The nuclear fuel element 500 may be cylindrically shaped, boxed shaped or the like.

The nuclear fuel element 500 may comprise a nuclear fuel 502 surrounded by cladding 504. A sensor 506 may be embedded within the nuclear fuel 502. The nuclear fuel 502 of the nuclear fuel element 500 may be formed, using the systems and methods of embodiments of the disclosure, to exhibit composition, chemical, or morphological (*e.g.*,
30 microstructural) differences in different regions along a height (*e.g.*, in the “Z” direction), and/or across a width (*e.g.*, in the “X” direction) thereof. In some embodiments, the differences may be in the form of gradients along the height and/or cross the width, or portions thereof. For example, the nuclear fuel 502 may be formed, using the systems and

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embodiments of the disclosure, to form regions of varying microstructures along a length and/or across a width thereof. Thus, the nuclear fuel element 500 may include a nuclear material (*e.g.*, a uranium-based nuclear material, such as a uranium-zirconium (UZr) material) with a porous microstructure in a porous zone 508, a less-porous/more-dense
5 microstructure in a less-porous zone 510, and a dense microstructure in a dense zone 512.

In some embodiments, the electrochemical cell 104 is used for fabricating uranium-zirconium fuel elements, such as the nuclear fuel element 500 of FIG. 5. Using the methods of embodiments of this disclosure, the nuclear fuel element 500 may be fabricated to include the dense zone 512 as a uranium-rich zone. Moreover, parasitic neutron-capturing elements,
10 such as a burnable absorber 514 (*e.g.*, poison material) may be embedded or distributed in the nuclear fuel 502. The cladding 504 may comprise stainless steel, and a barrier layer 516 (*e.g.*, of zirconium) may be provided between the nuclear fuel 502 and the cladding 504.

A nuclear fuel element such as the nuclear fuel element 500 may be additively manufactured, using any of the systems and methods described herein. For example, in some
15 embodiments, the nuclear fuel element 500 may be additively formed, through electrodeposition of the material of the nuclear fuel element 500, in layer-by-layer fashion in the z-direction. In some such embodiments the nuclear material of the less-porous zone 510, the porous zone 508, and the dense zone 512 may be electrodeposited in conjunction with one another, either also in conjunction with the material of the sensor 506 or with the sensor 506
20 inserted into the nuclear fuel 502 after the nuclear fuel 502 has been fabricated. The burnable absorber 514 may be inserted after or while fabricating the nuclear fuel 502. In some embodiments, the barrier layer 516 may be electrodeposited, in layer-by-layer fashion, along with the electrodeposition, in layer-by-layer fashion, of the nuclear fuel 502. Alternatively, after forming the nuclear fuel 502, it may be inserted within a tube comprising the barrier
25 layer 516 and the cladding 504.

The material of the nuclear fuel 502 may comprise aluminum-uranium alloys, uranium-zirconium alloys (*e.g.*, U-Zr, U-Pu-Zr) and/or may comprise oxide fuels (*e.g.*, UO₂, U₃O₈, and PuO₂-UO₂). Accordingly, the electrolytic bath 120 may include, but is not limited to, salts of uranium, aluminum, zirconium, cesium, plutonium, chlorine, and/or oxygen
30 dissolved therein, with the composition of the electrolytic bath 120 tailored according to the composition of the material to be electrodeposited. For example, in other embodiments, the electrolytic bath 120 may comprise radioactive elements for inclusion in the structure 108. In

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such embodiments, the electrolytic bath 120 may include hydrogen and/or carbon, as previously discussed.

Overall, by using an ionic bath for the electrolytic bath 120, the structure 108 (or structures 108), such as the structure of the nuclear fuel element 500 of FIG. 5, or the sub-
5 structures thereof, may be formed at relatively low temperatures compared to traditional manufacturing processes. Using a system disclosed herein (*e.g.*, system 100 of FIG. 1, system 200 of FIG. 2, system 300 of FIG. 3, system 400 of FIG. 4), the electrodeposition process may be conducted at relatively low temperatures, such as temperatures of 80°C or less, including room temperatures (*e.g.*, about 20°C to about 25°C). Moreover, any
10 radioactive materials to be electrodeposited by the system may be dissolved in the electrolytic bath 120. As a result, the radioactive materials may be highly confined and less susceptible to dispersion to the manufacturing atmosphere, compared to conventional powder deposition processes.

By way of example and not limitation, FIG. 6A and FIG. 6B are schematic
15 polarization curves for the electrodeposition of aluminum and for the electrodeposition of an aluminum-zirconium alloy, respectively. In an electrodeposition process, the potential applied and the current flowed by the controller (*e.g.*, controller 110 of FIG. 1 or first controller 204 of FIG. 2 or FIG. 3) to the electrochemical cell 104 may be varied, *e.g.*, during the electrodeposition, to selectively tailor one or more of the morphology (*e.g.*, shape,
20 microstructure, density) and/or composition of the material of the structure 108 formed on the substrate 106. As illustrated in FIG. 6A, in a system for electrodepositing aluminum (*e.g.*, the electrolytic bath 120 comprises aluminum ions), as the applied potential (*e.g.*, voltage) and current flow is reduced, the aluminum deposited may vary between a substantially fully dense deposit in the region E_a, a porous deposit in region E_m, and a microsphere or dendrite deposit
25 in region E_d. Therefore, the aluminum may be selectively deposited with a density/porosity gradient as the nozzle 122 is moved relative to the substrate 106, by controlling and adjusting the potential and/or current flow as the nozzle 122 is moved.

Similarly, the potential (*e.g.*, voltage) applied and/or current flowed by the controller (*e.g.*, the controller 110 of FIG. 1 or the first controller 204 of FIG. 2 or FIG. 3) to the
30 electrochemical cell 104 may be varied, *e.g.*, during the electrodeposition, to selectively tailor the relative composition of two or more elements being deposited from an electrolytic bath 120 by the system 100. As illustrated in FIG. 6B, for example, in a system with both zirconium and aluminum in the electrolytic bath 120, to deposit a structure 108 of an

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aluminum-zirconium alloy, the potential and current can be adjusted, *e.g.*, during the electrodeposition, to adjust the relative composition of zirconium to aluminum in the electrodeposited material. Notably, as the applied potential and current is reduced, the relative composition of zirconium and aluminum may be tuned by varying the concentration ratio of their precursors and deposition regions where their deposition reaction kinetics has different potential-dependence. In some embodiments, only zirconium may be deposited from zone 602, aluminum and zirconium may be co-deposited with a greater concentration of zirconium than aluminum from zone 604, and aluminum and zirconium may be co-deposited with a greater concentration of aluminum than zirconium from zone 606.

The fabrication (*e.g.*, flow) rate, or rate at which material may flow from the electrolytic bath 120, through the nozzle 122, to the substrate 106 (or the structure 108 thereon), may be varied by tailoring the size (*e.g.*, opening) of the nozzle 122 and by adjusting the kinetics of the reaction including, but not limited to, adjusting the temperature of the heater 124 and/or adjusting the potential or current applied by the controller 110 (FIG. 1) or the first controller 204 FIG. 2, FIG. 3. In some embodiments, the size of the nozzle 122 may be adjusted (*e.g.*, broadened or narrowed) during the electrodeposition by control via the controller 110 (or another controller of the system). For instance, physiochemical properties of the electrolytic bath 120 including, but not limited to, surface tension, viscosity, and diffusion coefficient are temperature dependent; accordingly, the process temperature may be varied, by controlling the heater 124, to selectively tailor the properties of the material deposited by the electrochemical processing unit (*e.g.*, the electrochemical processing unit 102 (FIG. 1), the electrochemical processing unit 202 (FIG. 2), the electrochemical processing unit 302 (FIG. 3)).

A method of forming a third-dimensional structure (*e.g.*, structure 108), which may be, for example, the nuclear fuel element 500 of FIG. 5, comprises providing the electrolytic bath 120 in the container 118 (*e.g.*, receptacle). The electrolytic bath 120 comprises a metal salt of a metal to be deposited. As previously discussed, counter electrode 114 (and, optionally, the reference electrode 116) may be at least partially disposed in the electrolytic bath 120 and may be coupled (*e.g.*, electrically coupled) to the substrate 106 disposed proximate the nozzle 122.

In operation, an electric current flow and/or a voltage difference may be applied between the substrate 106 (*e.g.*, the working electrode) and its corresponding counter electrode 114, resulting in the electrodeposition of a material (*e.g.*, a metal) derived from one

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or more salts (*e.g.*, metal salts) dissolved in the electrolytic bath 120. The voltage difference and/or current flow may be varied, *e.g.*, for and/or during deposition, to selectively tailor at least one of a morphology (*e.g.*, a microstructure, a density, a porosity) and/or a composition (*e.g.*, relative concentration of one element of an alloy to another element of the alloy) of the deposited material (*e.g.*, metal). In addition, the temperature of the heater 124 may be varied, *e.g.*, for and/or during deposition, to selectively tailor a physiochemical property of the metal salt as the metal salt flows through the nozzle 122. As illustrated in FIG. 4, the system may be configured with more than one electrochemical cell 104 to enable electrodeposition of more than one material concurrently.

10 In embodiments in which the structure 108 to be formed (*e.g.*, the nuclear fuel element 500 of FIG. 5) includes a nuclear material, a metal salt of a nuclear fuel metal may be dissolved in the electrolytic bath 120 and, during electrodeposition, may flow through the nozzle 122, as illustrated at arrow 704 of FIG. 7, and deposit on the surface of the substrate 106. As illustrated FIG. 4, the counter electrode 114 may comprise a material (A, B, or AB) similar to a material (A, B, or AB, respectively) of the structure 108 (*e.g.*, the nuclear fuel element 500 (FIG. 5)) being formed. The electrolytic bath 120 may also comprise a material, such as a salt (*e.g.*, A^{m+} , B^{n+} , or $A^{m+} + B^{n+}$) similar to the material (A, B, or AB, respectively) of counter electrode 114. For instance, each of the counter electrode 114, the electrolytic bath 120, and the structure 108 being formed may have at least one element in common.

20 As described with reference to FIG. 7, FIG. 8, and FIG. 9 below, a composition of the electrolytic bath 120 may be selected depending upon the operational mode--such as electrodeposition (for structure fabrication), electrodisolution (for structure machining), or electrochemical sensing (for structure or material analysis)--in which the system (*e.g.*, the system 100 of FIG. 1, the system 200 of FIG. 2, the system 300 of FIG. 3, the system 400 of FIG. 4) is to be used.

FIG. 7 illustrates an electrochemical cell 104, configured as an electrodeposition cell 700 for use in an electrodeposition process--using in any of the aforementioned systems--to form (*e.g.*, deposit, manufacture) the structure 108 (*e.g.*, the nuclear fuel element 500 of FIG. 5 or materials thereof) on the substrate 106, according to embodiments of the disclosure. In the electrodeposition process, the substrate 106 (*e.g.*, the working electrode) serves as a cathode and the counter electrode 114 serves as an anode. In the presence of a voltage (*e.g.*, potential difference) applied and a current flow (counter to electron flow as indicated by arrow "ne"), controlled by the controller 110 (FIG. 1) (or the first controller 204 of FIG. 2 or

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FIG. 3), ions (*e.g.*, metal salts) in the electrolytic bath 120 migrate from the electrolytic bath 120 in the container 118, through the nozzle 122, to the substrate 106 (*e.g.*, the working electrode), as indicated by dashed arrows 702. At the substrate 106 (*e.g.*, the working electrode), an electron-transfer reaction occurs to deposit the material of the structure 108. As
5 more and more material is deposited in this manner, the structure 108 increases in size. By moving the container 118 (and therefore also the nozzle 122) relative to the structure 108 (and therefore also the substrate 106), such as by operation of the electromechanical arm 126 (FIG. 1, FIG. 2, FIG. 3)--or, alternatively or additionally, by moving the substrate 106 (and therefore also the structure 108) relative to the container 118 (and therefore also the 122),
10 such as by operation of the XYZ platform 128 (FIG. 1, FIG. 2, FIG. 3)--the material of the structure 108 is deposited where additions to the structure 108 are desired, resulting in fabrication of a three-dimensional structure (*e.g.*, structure 108) on the substrate 106.

The electromechanical arm 126 and/or the XYZ platform 128 (FIG. 1, FIG. 2, FIG. 3) may manipulate the relative positions of the substrate 106 (and therefore the structure 108)
15 and the container 118 (and therefore the nozzle 122), such that the material may be selectively deposited and formed on the substrate 106, *e.g.*, layer-by-layer in the z-direction, the x-direction, and/or the y-direction. Therefore, the structure 108 may be formed to have complex shapes and/or dimensions. Moreover, the chemical composition and the morphology (*e.g.*, microstructure, density) of the material being formed can be adjusted,
20 during the fabrication process, by modifying the composition of the electrolytic bath 120, or the parameters of the electrodeposition therefrom (*e.g.*, current, voltage). Therefore, the systems of the disclosure are configured for selective modification of the composition and microstructure of the structure 108, including during the electrodeposition thereof.

Further, using multiple electrochemical cells 104 (*e.g.*, one or more of which may be
25 configured as the electrodeposition cell 700) and/or multiple nozzles 122 in the system, multiple different structures 108 and/or multiple different materials for the same structure 108 may be simultaneously or sequentially fabricated. Accordingly, the less-porous zone 510 of the nuclear fuel element 500 of FIG. 5 may be electrodeposited through one nozzle 122 (in communication with one electrolytic bath 120 of an electrochemical cell 104) while another
30 nozzle 122 (in communication with another electrolytic bath 120 of another electrochemical cell 104) electrodeposits the adjacent dense zone 512, in layer-by-layer fashion in the z-direction, before a third nozzle 122 (in communication with a third electrolytic bath 120 of a third electrochemical cell 104) electrodeposits the porous zone 508 on top of the less-porous

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zone 510, once the less-porous zone 510 has been fully electrodeposited. In another embodiment, the material of each of the porous zone 508, less-porous zone 510, and dense zone 512 may be deposited from the same electrolytic bath 120 and through the same or different nozzles 122, with the electrodeposition parameters adjusted, during the fabrication, to adjust the resulting porosity of the material being electrodeposited.

FIG. 8 illustrates an electrochemical cell 104, configured as an electrodisolution cell 800 for use in an electrodisolution process--using any of the aforementioned systems--to machine (*e.g.*, remove material from, dissolve material from) a structure 802 (*e.g.*, the structure 108 of FIG. 1 to FIG. 4 and FIG. 7, such as the nuclear fuel element 500 of FIG. 5 or materials thereof) on the substrate 106, according to embodiments of the disclosure. In some embodiments, the electrodisolution cell 800 may be used to machine (*e.g.*, remove material from, by electrodisolution) the substrate 106, itself. In the electrodisolution process, the substrate 106 (*e.g.*, the working electrode) serves as an anode and the counter electrode 114 serves as a cathode. In operation, a voltage difference (*e.g.*, potential difference) may be applied between the substrate 106 (*e.g.*, the working electrode) and the counter electrode 114, and current flows (counter to electron flow indicated by arrow "ne") from the substrate 106 (*e.g.*, the working electrode) toward the counter electrode 114. The voltage difference application and the current flows may be controlled by the controller 110 (FIG. 1) (or the first controller 204 of FIG. 2 or FIG. 3). In the presence of the applied voltage difference and the current flow, ions (*e.g.*, metal salts) migrate from the structure 802 through the nozzle 122, into and through the electrolytic bath 120, and then to the counter electrode 114, in the direction indicated by dashed arrows 806. At the counter electrode 114, the ions may react on the counter electrode 114 disposed in the electrolytic bath 120. Thus, material of the structure 802 is transferred to the counter electrode 114 by an electron-transfer reaction occurring at the counter electrode 114. Accordingly, a system (*e.g.*, the system 100 of FIG. 1, the system 200 of FIG. 2, the system 300 of FIG. 3, and/or the system 400 of FIG. 4) may, with one or more electrochemical cells 104 configured as the electrodisolution cell 800, be used for a machining (*e.g.*, etching) process to selectively remove material from the structure 802, which removal may form a recess 804 in the structure 802. As more and more material is dissolved and removed in this manner, the structure 802 decreases in size. By moving the container 118 (and therefore also the nozzle 122) relative to the structure 802 (and therefore also the substrate 106), such as by operation of the electromechanical arm 126 (FIG. 1, FIG. 2, FIG. 3)--or, alternatively or additionally, by moving the substrate 106 (and

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therefore also the structure 802) relative to the container 118 (and therefore also the nozzle 122), such as by operation of the XYZ platform 128 (FIG. 1, FIG. 2, FIG. 3)--the material of the structure 802 is removed (*e.g.*, dissolved, etched) from where material removal is desired, resulting in machining (*e.g.*, etching) from a three-dimensional structure (*e.g.*, structure 802) on the substrate 106.

The electromechanical arm 126 and/or the XYZ platform 128 (FIG. 1, FIG. 2, FIG. 3) may manipulate the relative positions of the substrate 106 (and therefore the structure 802) and the container 118 (and therefore the nozzle 122), such that the material may be selectively removed from the structure 802 in any of the z-direction, the x-direction, and/or the y-direction. Therefore, the structure 802 may be machined to have complex shapes and/or dimensions.

Further, using multiple electrochemical cells 104 (*e.g.*, one or more of which may be configured as the electrodisolution cell 800) and/or multiple nozzles 122 in the system, multiple different structures 108 and/or multiple different materials from the same structure 108 may be simultaneously or sequentially machined (*e.g.*, etched). In another embodiment, material from each of multiple regions of the structure 802 may be etched using the same electrolytic bath 120 and counter electrode 114 and through the same or different nozzles 122.

Accordingly the systems embodiments of the disclosure may be used for machining (*e.g.*, etching) a structure by an electrodisolution process with selective removal of material from the structure. Such a system reduces waste compared to, for example, systems for conventional leaching process. Notably, the electrodisolution material-removal processes of embodiments of the disclosure may effectively recycle and collect previously-deposited material. That is, in some embodiments the system may have been used with the electrochemical cell 104 configured as the electrodeposition cell 700 of FIG. 7, to form the structure 108 by depositing material from the counter electrode 114, through the electrolytic bath 120 with its ionic liquid (IL) and ions (*e.g.*, metal salts), through the nozzle 122 and onto the substrate 106 to form the structure 108. The structure 108 may, thereafter, serve as the structure 802 of FIG. 8, and the same electrochemical cell 104 (or another electrochemical cell 104) may then be configured as the electrodisolution cell 800 of FIG. 8, *e.g.*, by adjusting the direction of the current flow, *etc.*, to selectively remove (*e.g.*, dissolve) material from the structure 802. The dissolved material may travel back through the nozzle 122, into the electrolytic bath 120 with the ionic liquid (IL) and ions (*e.g.*, metal salts), and to the

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counter electrode 114 where the material redeposits. In some embodiments, the electrodeposition and the electrodisolution processes, illustrated in FIG. 7 and FIG. 8, respectively, may be conducted simultaneously or concurrently and using one or more of the electrochemical cells 104 discussed above.

5 For both the electrodeposition (FIG. 7) and the electrodisolution (FIG. 8), the counter electrode 114 may comprise a material (M) similar to a material (M) of the structure 108 (FIG. 7) or structure 802 (FIG. 8). The electrolytic bath 120, for both the electrodeposition and the electrodisolution, may comprise a material (*e.g.*, a salt) (M^{n+}) similar to the material (M) of the 114. For instance, each of the 114, the 120, and the structure 108 (or the structure 10
802) may have at least one element in common.

FIG. 9 illustrates an electrochemical cell 104, configured as an electrochemical-sensor cell 900 for use in an electrochemical sensing process--using any of the aforementioned systems--to analyze or otherwise determine a composition or characteristic of the structure 108 or an analyte in a gas, liquid, or solid phase of the structure 108, according to
15 embodiments of the disclosure. In particular, a composition of the structure 108 may be determined (*e.g.*, analyzed, electrochemically sensed) by measuring a potential difference between the substrate 106 (*e.g.*, the working electrode) and the counter electrode 114. The measurements of the potential difference may be taken by, *e.g.*, one or more of the controllers of the system. The measured potential difference may be correlated to a difference in
20 composition of the counter electrode 114 and the structure 108 to determine the composition of the structure 108 or other item subject to analysis. Alternatively, the potential difference between the substrate 106 (*e.g.*, the working electrode) and the reference electrode 116, if the latter is included, may be measured to determine the composition of the structure 108. Similarly, a current flow (*e.g.*, dashed arrow 902) (*e.g.*, counter to the direction of electron
25 flow, as indicated by arrows "ne") may be controlled and measured as part of the electrochemical sensing. The current and voltage relationship, measured using a system that includes an electrochemical cell 104 configured as the electrochemical-sensor cell 900, may be unique to the materials present (material A of the structure 108, material B dissolved in the ionic liquid (IL) of the electrolytic bath 120, and/or material C of the counter electrode 114).
30 Therefore, by measuring the current and voltage using the system and the electrochemical-sensor cell 900, the composition or characteristics of the materials may be discernable.

While the electrodeposition (*e.g.*, of FIG. 7), the electrodisolution (*e.g.*, of FIG. 8), and the electrochemical sensing (*e.g.*, of FIG. 9) have been described, herein, as separate

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processes, the disclosure is not so limited. In some embodiments, for example, one or more of the foregoing processes may be conducted sequentially (*e.g.*, consecutively) or concurrently. And, any electrochemical cell 104 may, in some embodiments, be reconfigured for another of the foregoing processes, *e.g.*, by adjusting control parameters, voltage amounts
5 or directions, and/or current amounts and direction.

The embodiments of the disclosure are not limited to electrochemical cells 104 of a shape and structure illustrated in the figures. In other embodiments, for example, one or more of the electrochemical cells 104 of a system may be configured as syringes, with the body of the syringe providing the container 118 of the electrochemical cell 104, and the liquid
10 contents of the syringe being formulated as the electrolytic bath 120. The rate of dispensation of the electrolytic bath 120 from a syringe-type electrochemical cell 104, or the retraction of material back up into the syringe-type electrochemical cell 104, may be controlled by controlling the rate of engagement of a plunger of the syringe, which rate of engagement may be controlled by a controller of the system (*e.g.*, any of the aforementioned controllers or
15 another controller).

While the system (*e.g.*, system 100 of FIG. 1, system 200 of FIG. 2, system 300 of FIG. 3, system 400 of FIG. 4) and methods have been described with respect to formation of a nuclear fuel element (*e.g.*, the nuclear fuel element 500 of FIG. 5), the present disclosure is not so limited. Any of the systems and/or methods may be used to manufacture (*e.g.*, by
20 electrodeposition), machine (*e.g.*, by electrodisolution), and/or analyze (*e.g.*, by electrochemical sensing) metallic and non-metallic structures for nuclear and non-nuclear applications. For instance, any of the foregoing systems and foregoing electrochemical cells 104 may also have non-nuclear applications, including the manufacturing of any object in place of an additive manufacturing or advanced manufacturing process, including, but not
25 limited to, the manufacture of batteries, solar cells, electronics, propellers, turbine blades, and sensors for process control and environmental monitoring.

While embodiments of the disclosure may be susceptible to various modifications and alternative forms, specific have been described in detail herein. However, it should be understood that the disclosure is not limited to the particular forms disclosed. Rather, the
30 disclosure encompasses all modifications, variations, combinations, and alternatives falling within the scope of the disclosure as defined by the following appended claims and their legal equivalents.

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CLAIMS

What is claimed is:

- 5 1. An electrochemical system comprising:
at least one electrochemical cell comprising:
a receptacle containing an electrolytic bath;
at least one nozzle opening from the receptacle toward and proximate a
substrate configured as a working electrode of the at least one
10 electrochemical cell; and
a counter electrode disposed in the electrolytic bath,
the at least one electrochemical cell selectively configurable between a configuration
for electrodeposition of material to form a structure on the substrate and a
configuration for electrodisolution of material from the structure on the
15 substrate.
2. The electrochemical system of claim 1, further comprising at least one of an
electromechanical arm and an XYZ platform configured to control relative movement
between the at least one nozzle and the substrate.
20
3. The electrochemical system of claim 1, further comprising at least one
controller configured to apply a current or voltage to the counter electrode and the substrate
configured as the working electrode.
- 25 4. The electrochemical system of claim 1, wherein the electrolytic bath
comprises an ionic liquid and at least one a nuclear fuel material salt dissolved in the ionic
liquid.
- 30 5. The electrochemical system of claim 1, wherein the electrolytic bath
comprises an ionic liquid and at least one of a uranium salt and a zirconium salt dissolved in
the ionic liquid.

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6. The electrochemical system of claim 1, further comprising a heater disposed about the at least one nozzle, the substrate, or both.

7. The electrochemical system of claim 1, wherein each of the receptacle and the
5 substrate are movable in three dimensions.

8. The electrochemical system of claim 1, wherein the electrochemical system comprises a plurality of the electrochemical cells, each electrolytic bath of the electrochemical cells having a different chemical composition.

10

9. The electrochemical system of claim 1, wherein the electrochemical system further comprises a plurality of controllers, the plurality of controllers comprising:

at least one controller configured to control a voltage difference and a current flow between the counter electrode and the substrate configured as the working
15 electrode; and

at least one other controller configured to control movement of the at least one nozzle over the substrate.

10. The electrochemical system of any one of claims 1 through 9, wherein the at
20 least one nozzle is movable in three dimensions relative to the substrate.

11. A method of using an electrochemical cell, comprising:

providing an electrolytic bath in a receptacle, the electrolytic bath comprising a metal salt;

25 disposing a counter electrode at least partially within the electrolytic bath;

coupling the counter electrode to a working electrode; and

in the presence of at least one of a voltage difference and a current flow between the working electrode and the counter electrode, flowing the metal salt through a nozzle coupled to the receptacle.

30

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12. The method of claim 11, wherein flowing the metal salt through the nozzle coupled to the receptacle comprises flowing the metal salt from the electrolytic bath, through the nozzle, and to the working electrode to form a structure comprising a metal of the metal salt, wherein the working electrode serves as a cathode, and wherein the counter electrode
5 serves as an anode.

13. The method of claim 11, wherein flowing the metal salt through the nozzle coupled to the receptacle comprises flowing the metal salt from a structure on the working electrode, through the nozzle, through the electrolytic bath, and to the counter electrode, to
10 form a recess in the structure on the working electrode, wherein the working electrode serves as an anode, and wherein the counter electrode serves as a cathode.

14. The method of claim 11, wherein flowing the metal salt through the nozzle comprises:
15 flowing the metal salt through the nozzle toward the working electrode to electrodeposit a metal on the working electrode; and thereafter, flowing the metal salt from the metal on the working electrode through the nozzle to deposit the metal on the counter electrode within the receptacle.

20 15. The method of claim 11, further comprising measuring at least one of the voltage difference and the current flow to identify a composition of at least one of the counter electrode, the electrolytic bath, and a structure on the working electrode.

16. The method of any one of claims 11 through 15, wherein the metal salt
25 comprises uranium or plutonium.

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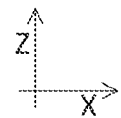
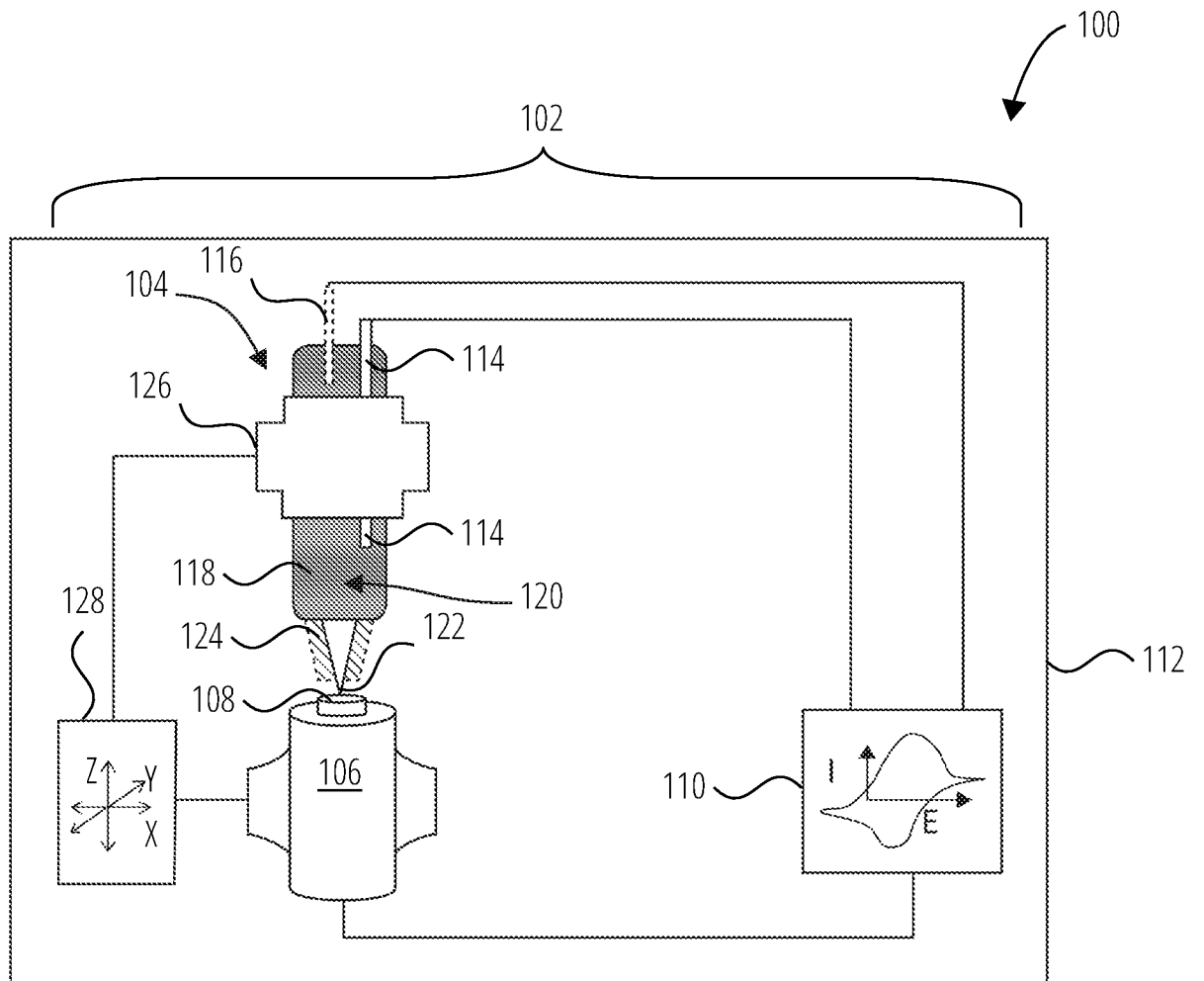


FIG. 1

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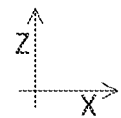
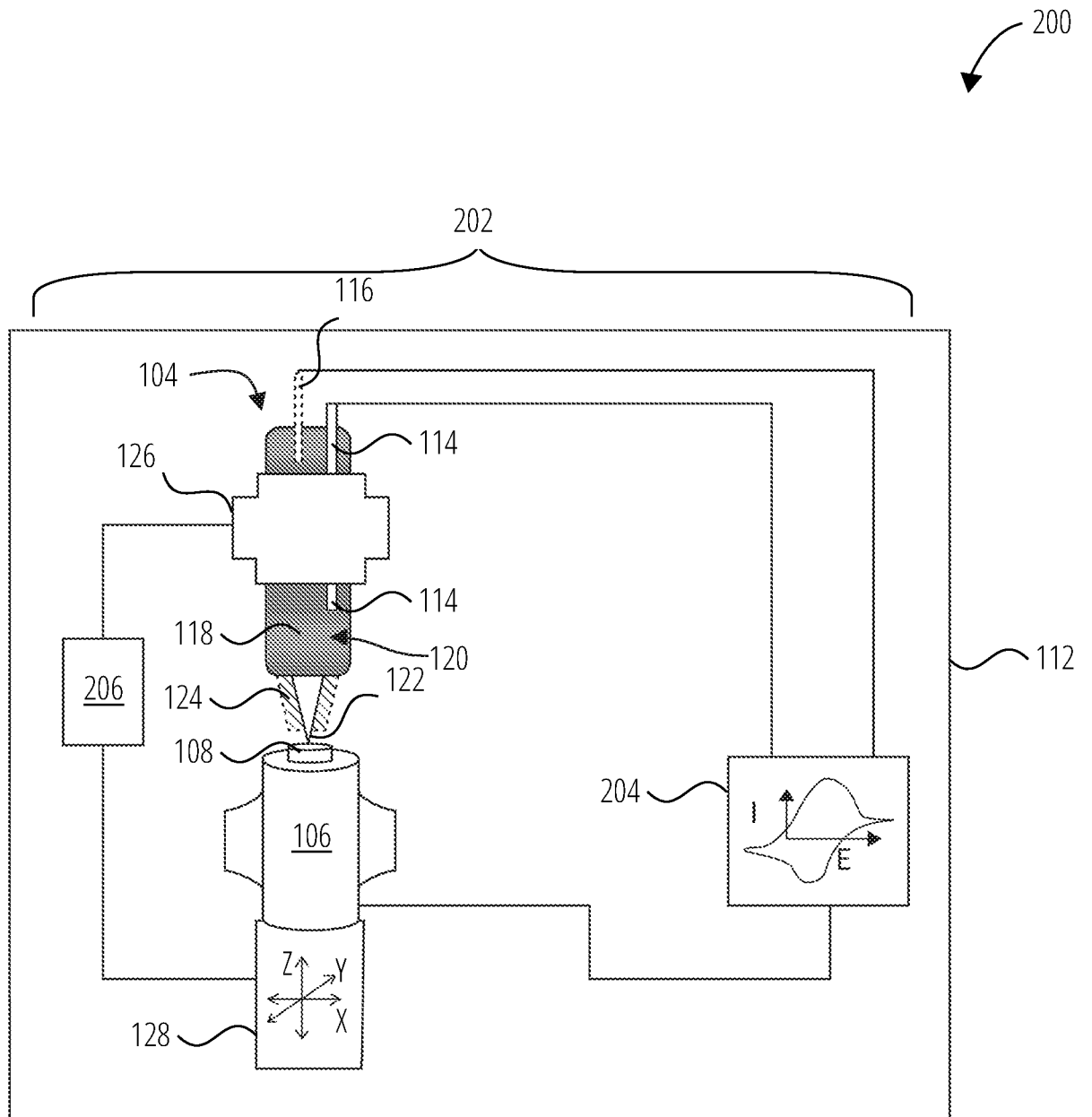


FIG. 2

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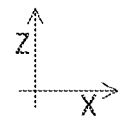
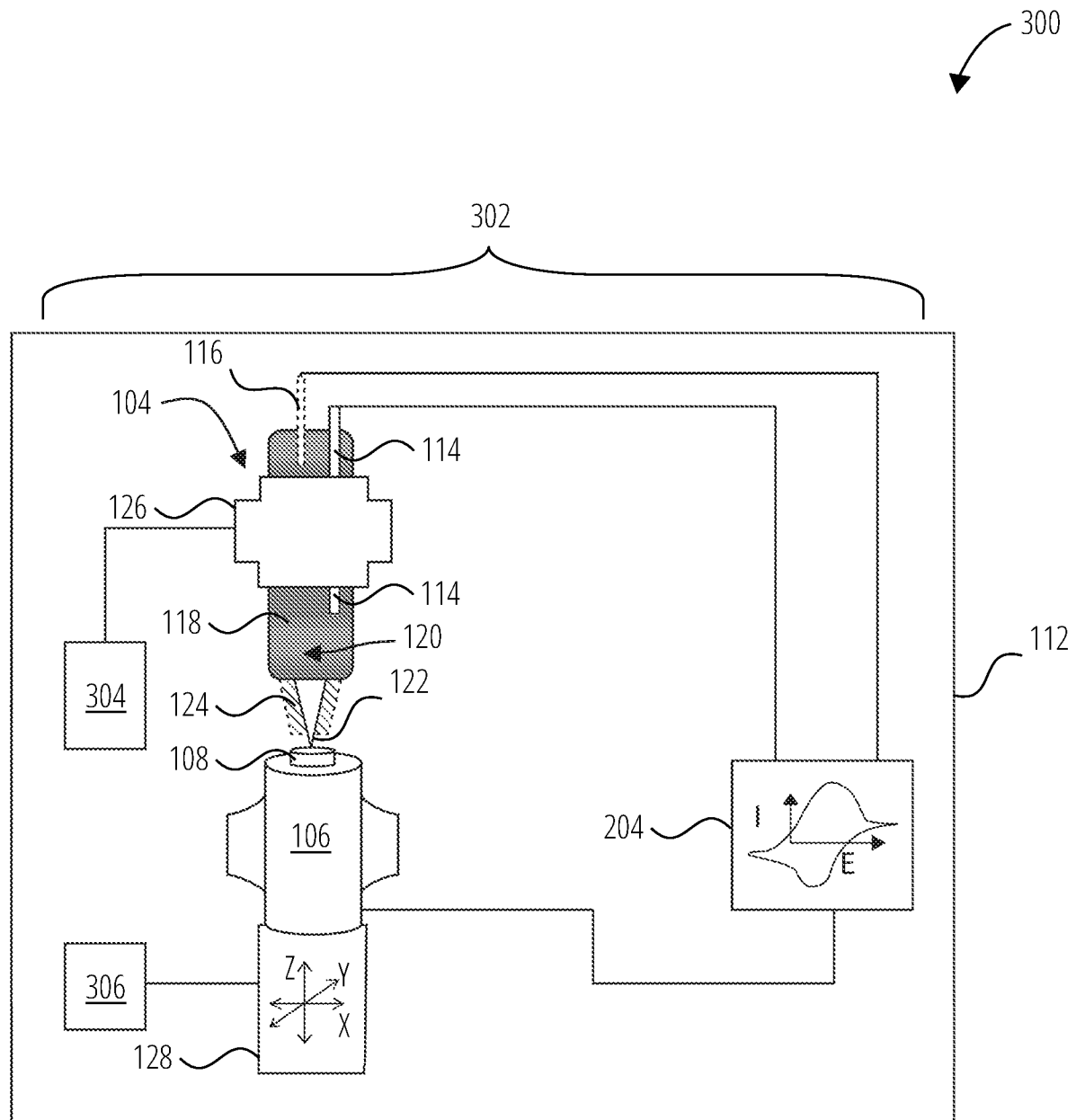


FIG. 3

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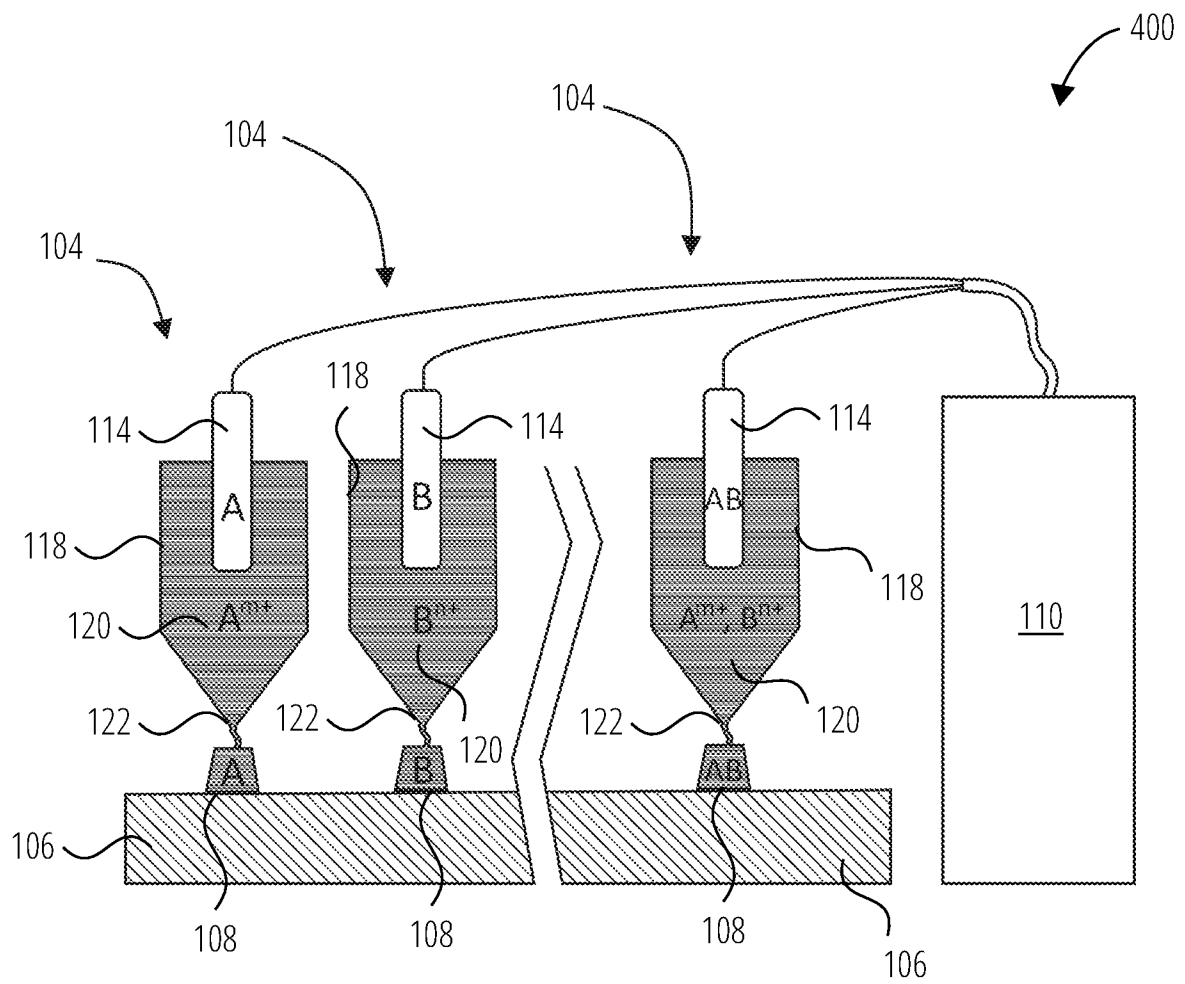


FIG. 4

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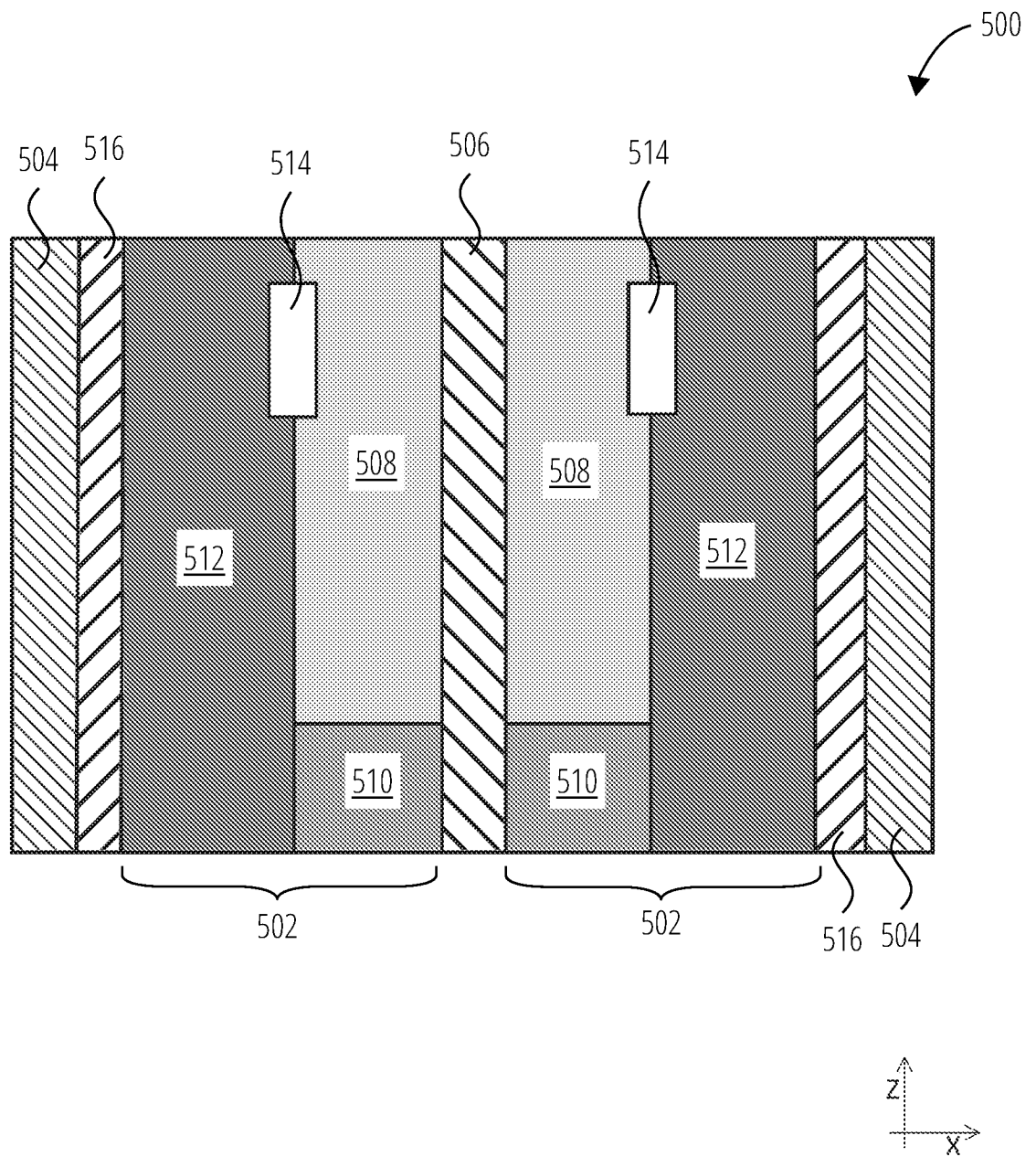


FIG. 5

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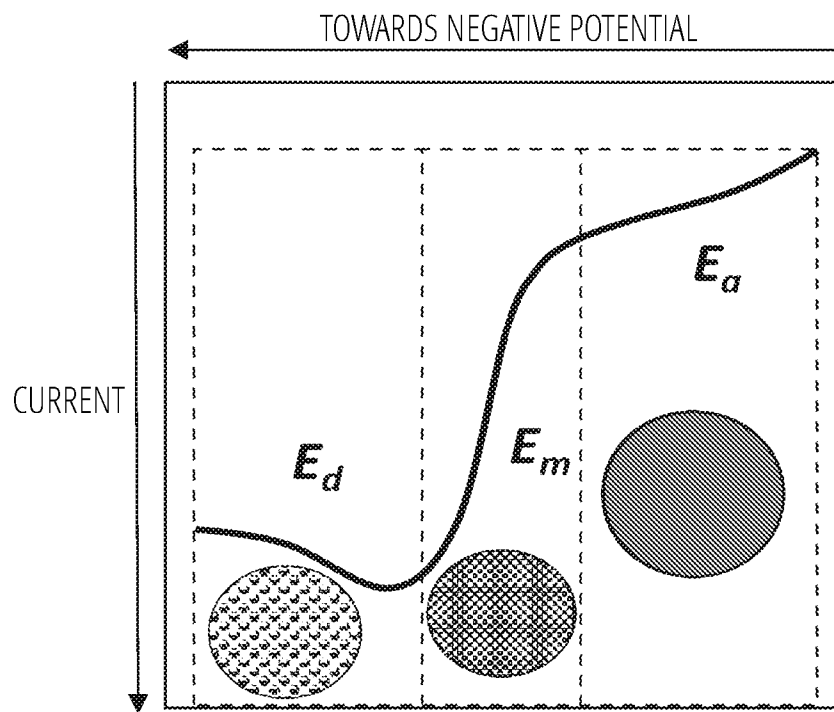


FIG. 6A

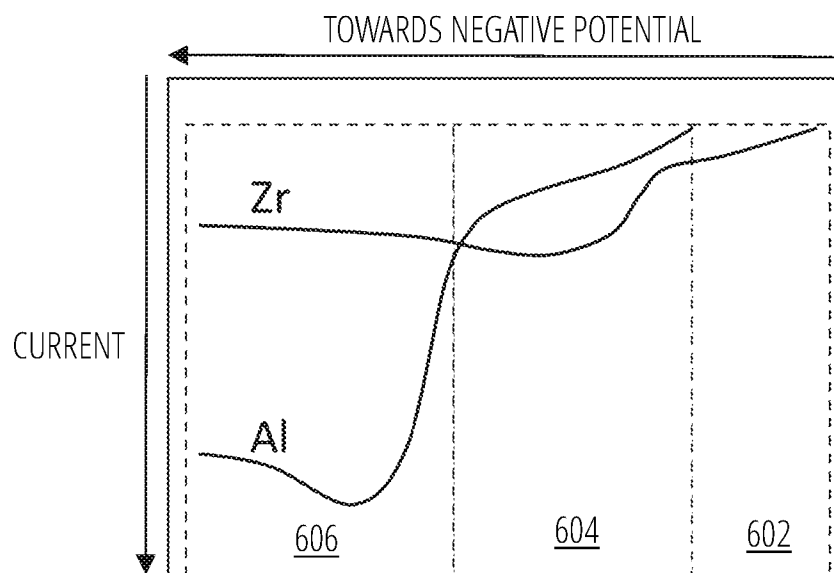


FIG. 6B

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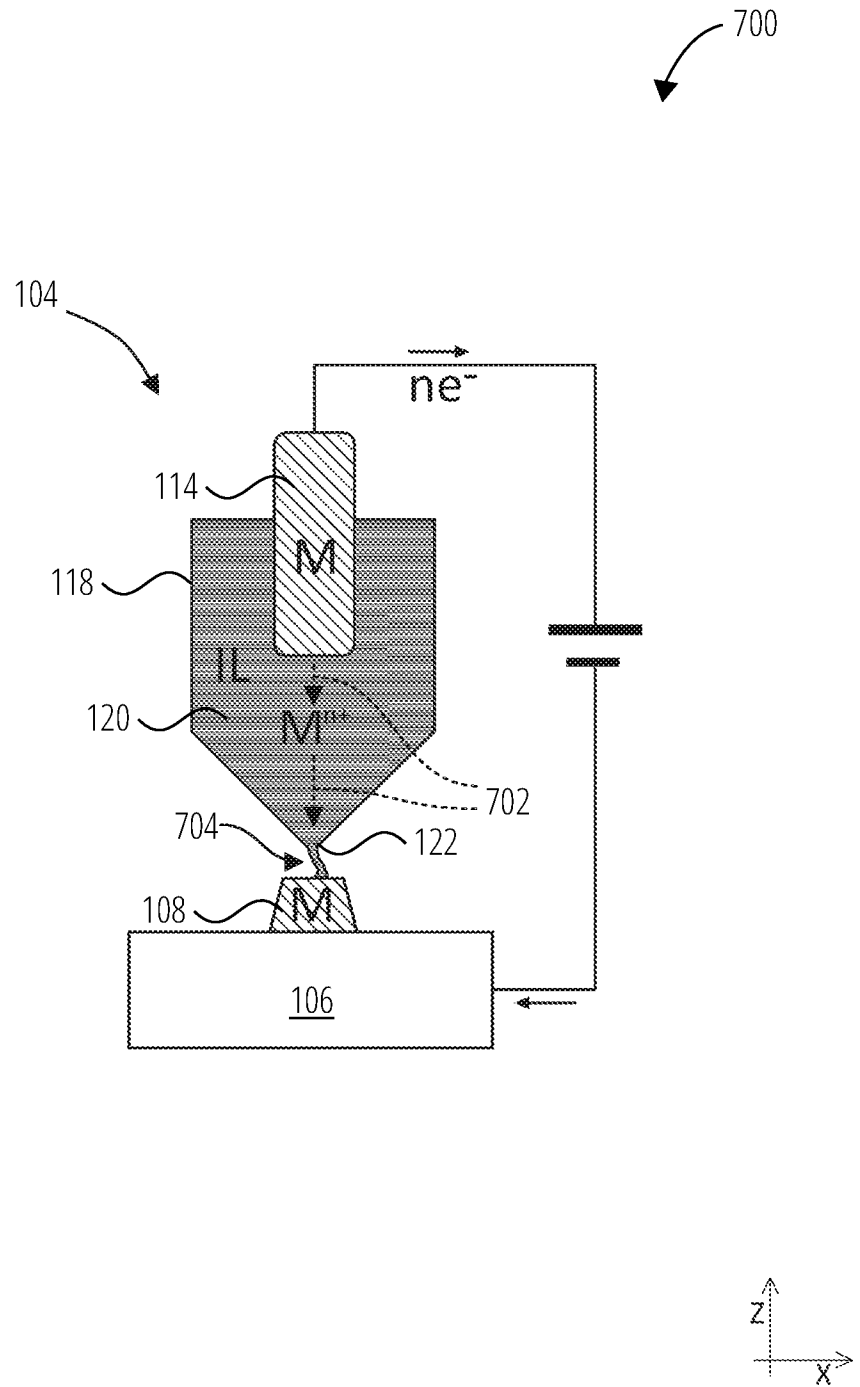


FIG. 7

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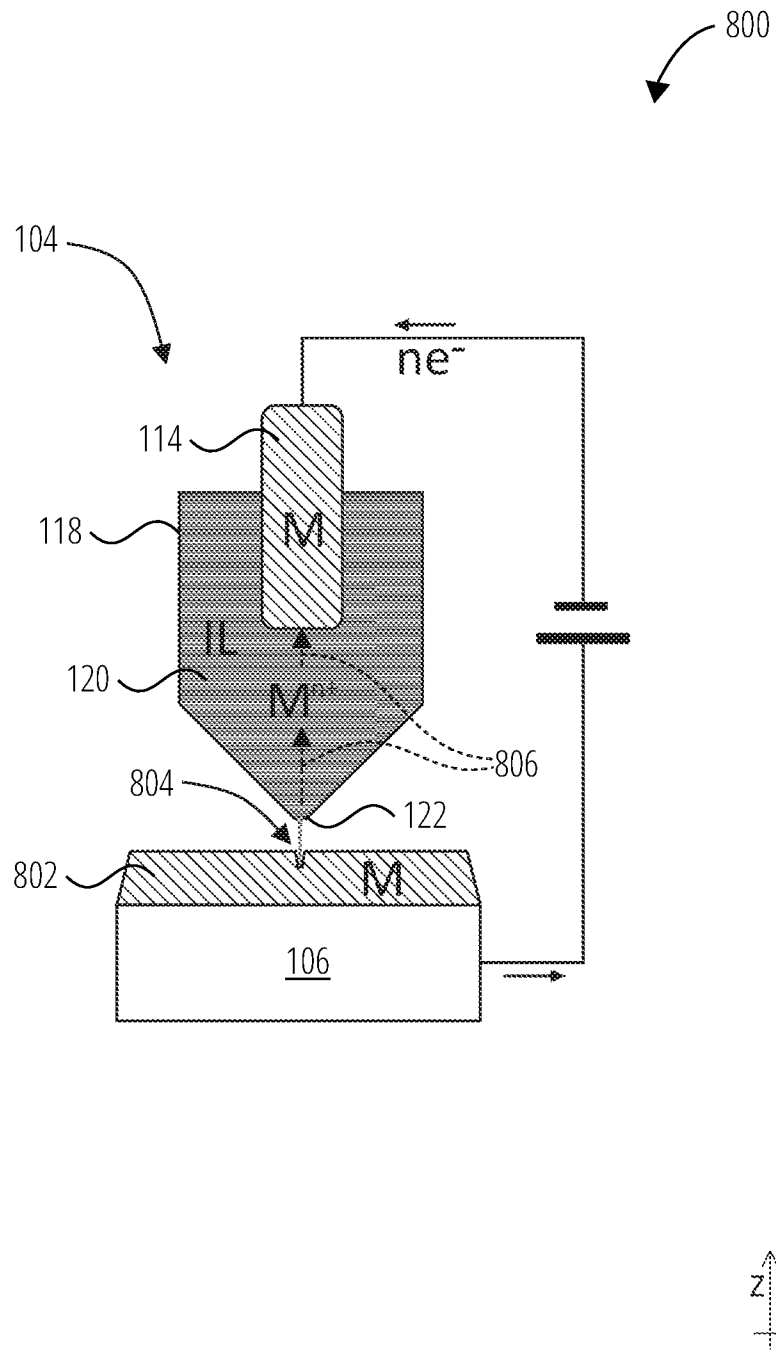


FIG. 8

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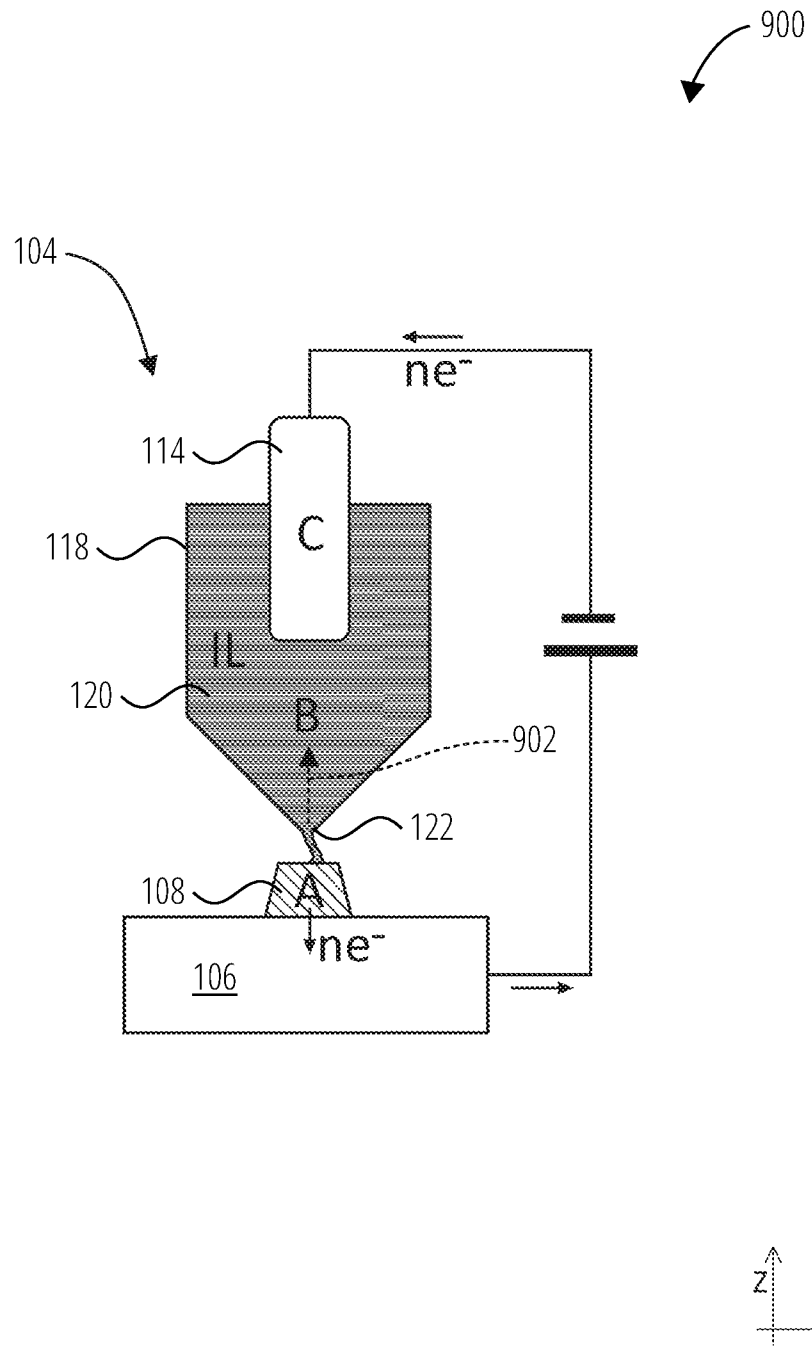


FIG. 9

INTERNATIONAL SEARCH REPORT

International application No.

PCT/US 19/65399

A. CLASSIFICATION OF SUBJECT MATTER

IPC - H01M 4/04, H01M 6/40; C23C 18/54; C25D 1/00, C25D 3/02 (2020.01)

CPC - H01M 4/04, 4/0402, 4/0438, 4/045, 4/0452, 4/0454, 4/0438, 6/40; C23C 18/54; C25D 1/003, 3/02; H01L 21/283, 21/288, 21/2885, 24/27; B29C 48/30, 64/00, 64/10, 64/106, 64/20, 64/209; B33Y 10/00, 30/00, 80/00; B22F 3/115; H05K 3/12; Y10T 29/49108

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)

See Search History document

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

See Search History document

Electronic data base consulted during the international search (name of data base and, where practicable, search terms used)

See Search History document

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category*	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X --- Y	WO 2017/055338 A1 (ETH ZURICH) 06 April 2017 (06.04.2017) entire document, Fig 1, 3, 4; pg 1, ln 5-6; pg 4, ln 20-21; pg6, ln 2-4, 8-13; pg 8, ln 8-11, pg9, ln 4-8, 11-14; pg 10, ln 14-22	11 ----- 1-10, 12-16
Y	US 8,147,660 B1 (MAYER) 03 April 2012 (03.04.2012) entire document, Fig 4; Col 6, ln 54-65; Col 8, ln 57-67; Col 9, ln 1-3; Col 21, ln 65-67; Col 22, ln 1-15; Col 30, ln 30-33	1-10, 12-14, 16/(12-14)
Y	US 2012/0090987 A1 (GORER et al.) 19 April 2012 (19.04.2012), para [0015], [0032]	8, 10/(8)
Y	US 2013/0142566 A1 (YU) 06 June 2013 (06.06.2013), para [0099], [0101], [0105], [0127]	6, 7, 9, 10/(6-7,9), 15, 16/(15)
Y	WO 01/13379 A1 (BRITISH NUCLEAR FUELS PLC) 22 February 2001 (22.02.2011), para [0013], [0014], [0023]	4, 5, 10/(4-5), 16/(11-15)

☐ Further documents are listed in the continuation of Box C.☐ See patent family annex.

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"E" earlier application or patent but published on or after the international filing date

"L" document which may throw doubts on priority claim(s) or which is cited to establish the publication date of another citation or other special reason (as specified)

"O" document referring to an oral disclosure, use, exhibition or other means

"P" document published prior to the international filing date but later than the priority date claimed

"T" later document published after the international filing date or priority date and not in conflict with the application but cited to understand the principle or theory underlying the invention

"X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone

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Date of the actual completion of the international search

01 March 2020

Date of mailing of the international search report

23 MAR 2020

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