

[54] CLOSED-CYCLE ELECTROCALORIC REFRIGERATOR AND METHOD

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[58] Field of Search 62/3

[56] References Cited

UNITED STATES PATENTS

Table with 3 columns: Patent Number, Date, Inventor Name, and Reference Number. Includes entries for Chilowsky, Garwin, Pearl, and Otter.

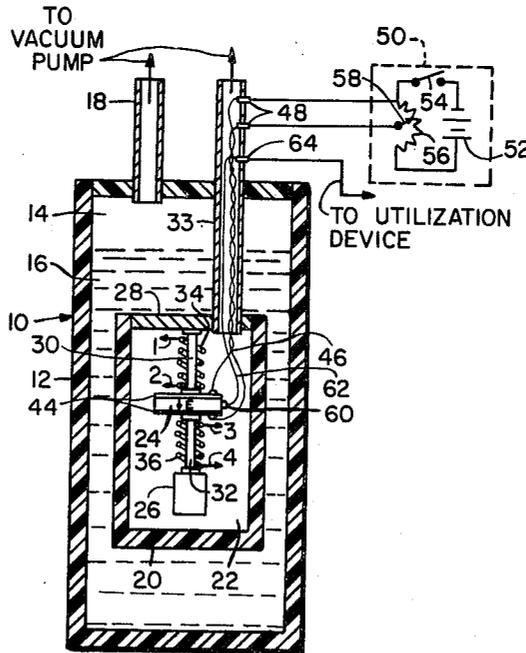
3,436,924 4/1969 Lawless 62/3

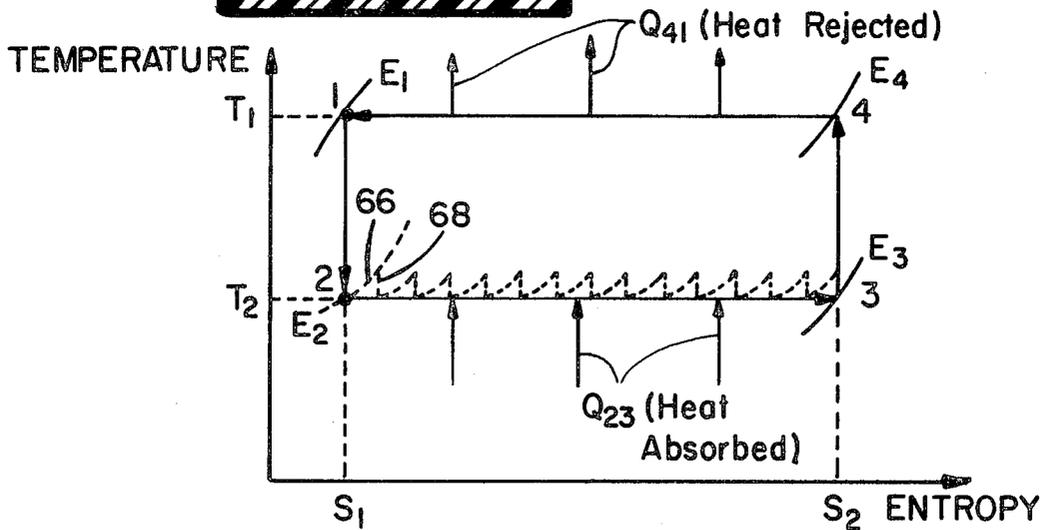
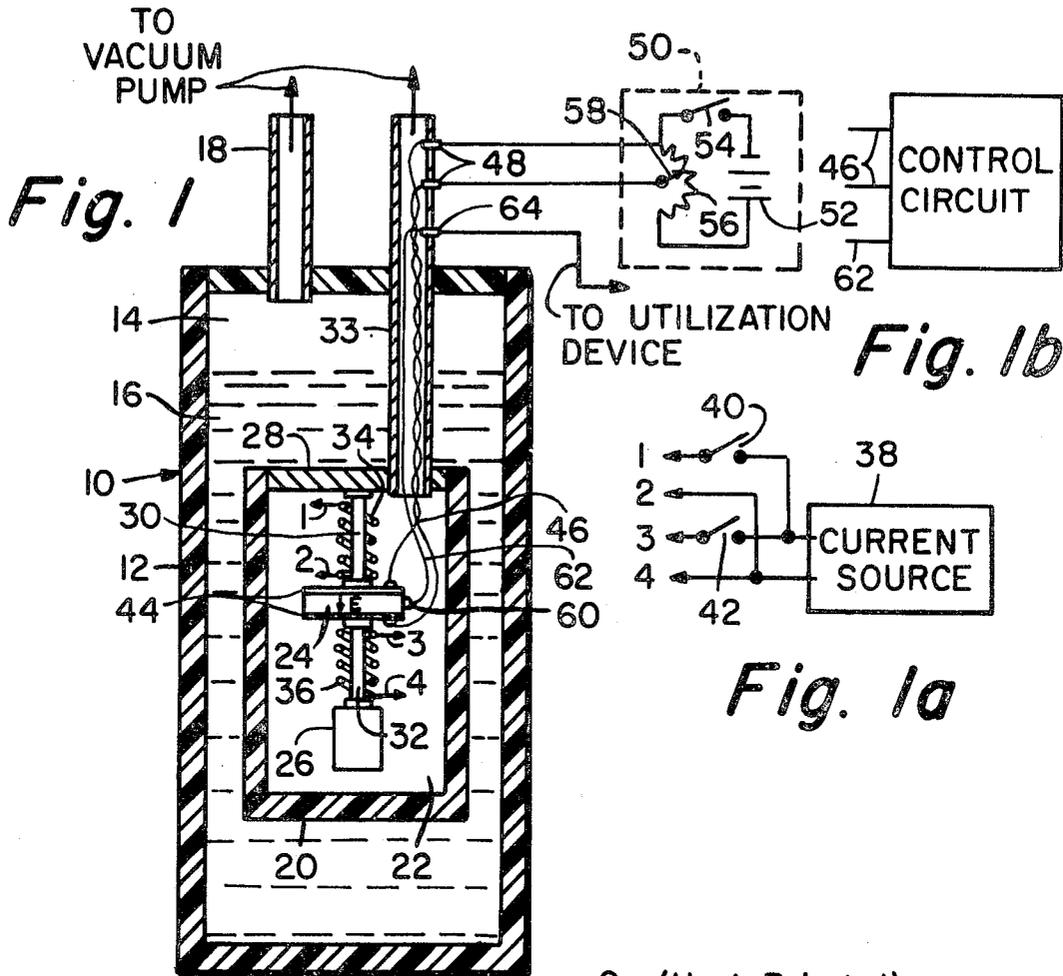
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[57] ABSTRACT

A closed-cycle solid state refrigerator is disclosed that is capable of accurately and efficiently achieving extremely low temperatures. A second stage of the refrigerator comprises an evacuated chamber in which is disposed a solid state refrigerant that is connected by a first thermal valve to a first stage such as a cryostat and by a second thermal valve to a heat generating load that is to be cooled.

22 Claims, 4 Drawing Figures





**Fig. 2**

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## CLOSED-CYCLE ELECTROCALORIC REFRIGERATOR AND METHOD

### BACKGROUND OF THE INVENTION

The development in recent years of devices and systems which function at cryogenic temperatures, such as cryogenic computer memories, high efficiency gallium arsenide diodes, high-field superconducting alloys, masers, and the like has led to a search for more economical and less complex means for obtaining cryogenic environments. To achieve temperatures on the order of 1° K., expensive cryogenic pumping systems using helium as a refrigerant have been required.

In some cases several types of helium cooling systems have been combined to achieve extremely low temperatures. For example, the conventional helium cryostat cooling system has been used to produce temperatures on the order of 4° K., at which point the well-known helium pump has been employed to produce additional cooling to temperatures as low as about 1° K. Another helium apparatus, commonly known in the art as an Oxford instrument, has also been used in combination with the helium pump system to provide cryogenic environments as low as 0.02° K.

Elaborate helium cooling systems have also been used to cool paramagnetic refrigerant materials to intermediate temperatures sufficient to initiate paramagnetic cooling processes. Though temperatures as low as about 0.001° K. have been achieved by paramagnetic cooling methods, the maximum starting temperature from which effective paramagnetic cooling can be initiated is only about 1.2° K. Such low starting temperatures are not obtainable by means other than complex helium refrigeration systems.

In my U.S. Pat. No. 3,436,924 entitled "Pараelectric Refrigeration Method and Apparatus" there is disclosed a paraelectric refrigerator which utilizes a solid-state refrigerant that cools paraelectrically from a starting temperature as high as about 20° K. Electrodes are disposed on opposite sides of the paraelectric refrigerant, and a load disposed adjacent one of the electrodes is in thermal contact with the refrigerant. A cryogenic refrigerant is introduced into a chamber containing the above-described assembly which is thereby cooled to within the paraelectric cooling range of the paraelectric refrigerant which is then isothermally polarized by an applied electric field. The refrigerant and load are then thermally isolated by evacuating the chamber and the refrigerant is caused to cool paraelectrically by adiabatically depolarizing the paraelectric refrigerant by removing the applied electric field. Although the system of said patent is capable of providing paraelectric cooling from a starting temperature as high as about 20° K., the efficiency of operation of said system is low and the final temperature to which the refrigerant and load are ultimately cooled is not easily controlled.

### SUMMARY OF THE INVENTION

It is therefore an object of the present invention to provide a refrigerator capable of accurately refrigerating a load at a preselected temperature.

Another object of the present invention is to provide a refrigerator that is capable of cooling in a manner which is more efficient than has been heretofore possible.

Another object of the present invention is to provide a refrigeration method for accurately and efficiently refrigerating a load.

Briefly, the present invention relates to a closed-cycle refrigerator for refrigerating a mass of material. This refrigerator includes an electrocaloric refrigerant and a sink for absorbing heat from the refrigerant at a substantially constant temperature within the electrocaloric refrigeration range of the refrigerant. Valve means are provided for thermally connecting and separating the refrigerant and the material to be cooled and also for thermally connecting and separating the refrigerant and sink. Means are provided for isolating the refrigerant and the material, and means are provided for electrically polarizing and depolarizing the refrigerant.

This refrigerator functions in accordance with the following process. The refrigerant and material to be refrigerated are thermally connected to the heat sink, the temperature of which is  $T_1$ , so that the refrigerant and the material are substantially cooled to temperature  $T_1$ . Then, after thermally isolating the material, the polarization of the refrigerant is isothermally changed toward a first state of polarization by applying thereto an electric field of intensity  $E_1$  to reduce the entropy of the refrigerant to a value  $S_1$ . The refrigerant is then thermally isolated, and the polarization thereof is adiabatically changed toward a second state of polarization opposite the first state by changing the intensity of the field from intensity  $E_1$  to an intensity  $E_2$  to reduce the temperature of the refrigerant from temperature  $T_1$  to a temperature  $T_2$ . The refrigerant is then thermally connected to the material to decrease the temperature thereof below the temperature  $T_1$ .

### BRIEF DESCRIPTION OF THE DRAWINGS

FIGS. 1 and 1a show an elevation view of a closed-cycle electrocaloric refrigeration system with associated control circuitry illustrating one embodiment of the present invention.

FIG. 1b illustrates an alternate embodiment of the control circuitry of the present invention.

FIG. 2 is a graph of the temperature vs. entropy characteristics of a solid-state material employed as a refrigerant in the system of FIGS. 1 and 1a during a typical Carnot-cycle operating sequence.

### DETAILED DESCRIPTION

The solid state, closed-cycle refrigerator of the present invention is capable of cooling by either of two operating cycles depending upon whether the electrocaloric material utilized therein cools upon the adiabatic depolarization thereof (type I materials or paraelectric materials) or the adiabatic polarization thereof (type II materials). Since the type I materials are more conventional, the operation of the invention will first be described with reference to such type I or paraelectric materials.

As used herein, an electrocaloric refrigerant is a material the temperature of which changes, under adiabatic conditions, in response to the application of an electric field. This means that the entropy of the material is controlled externally by polarization and depolarization of the material. An electrocaloric refrigerator is defined as one which employs an electrocaloric refrigerant. As used herein, the electrocaloric refrigeration range is that range of temperatures for which entropy of an electrocaloric material is substantially reduced by changing the polarization of the material by an applied electric field. This means that the amount of entropy decrease should be about as large as the total entropy in the absence of the polarizing field. The electrocaloric starting temperature is defined as any temperature within the electrocaloric refrigeration range.

Referring to FIG. 1 there is shown an example of the closed-cycle electrocaloric refrigeration system of the present invention. A first stage of two refrigeration stages of the system consists of a conventional helium cryostat 10 which is employed to provide the necessary electrocaloric starting temperature, and to act as a high temperature isothermal heat sink during closed-cycle operation. Cryostat 10 consists of a pressure vessel 12 having a vacuum chamber 14 which, under the usual operating conditions, contains a quantity of liquid helium 16. A pipe 18 communicating with chamber 14 may be employed to introduce helium into the chamber, and by connecting it to a suitable vacuum pump, not shown, provides means whereby a partial vacuum may be drawn on chamber 14 if desired. Any suitable and well-known type of cryogenic refrigerator, having the capacity to attain the desired electrocaloric starting temperature, may be used in place of cryostat 10. Examples of suitable substitutes include the well-known Stirling-cycle refrigerator, the Gifford-McMahon cycle refrigerator, a helium dilution refrigerator, and the like.

The second stage of the system is an electrocaloric refrigeration stage consisting of pressure vessel 20 having an overall size suitable for suspension within chamber 14. Under operating conditions of the first stage, it is preferable that vessel 20 be completely immersed in the liquid helium 16 as shown. Vessel 20 also has a vacuum chamber 22 in which is suspended any suitable electrocaloric refrigerant 24, and a mass of any material, hereinafter referred to as a thermal load 26, which material is to be refrigerated by the system in accordance with the process of the present invention. Note that the load is a source of heat, and this heat is to be absorbed so as to keep the load temperature constant. Refrigerant 24 is suspended from a thermally conductive lid 28, so as not to contact the defining walls of the chamber 22, by means of a conventional superconductive thermal switch or valve 30. The load 26 is similarly suspended from the refrigerant 24 by means of a second thermal valve 32, preferably of the same type as valve 30. A pipe 33, communicating with chamber 22, can be connected to a suitable vacuum pump, not shown, in order that a high quality vacuum, preferably on the order of about 1 micron of Hg or less, can be drawn on chamber 22 during the operating sequence of the system. Such a vacuum level insures that the materials within chamber 22 can be effectively thermally isolated from chamber 14 when the refrigeration process of the present invention so requires. A lower quality vacuum may be employed successfully, however, the quality of thermal isolation obtainable, and hence, the refrigeration efficiency of the system may be expected to suffer as a result. The load and refrigerant could also be thermally isolated by filling the space between these elements and the inner walls of the chamber with an insulating material.

Those skilled in the art will appreciate that the superconductive thermal valves 30 and 32 undergo a transition to a superconducting state at a suitable low temperature, but the application of a magnetic field can drive this type of material back into the nonsuperconducting or normal state. Since the superconducting state has a very low thermal conductivity relative to the normal state, a thermal valve property is produced. Magnetic coils 34 and 36 surround the valves 30 and 32, respectively.

Referring additionally to FIG. 1a, an electric current source 38 is connected through switch 40 across coil 34 as indicated by match lines labeled 1 and 2, and through a switch 42 across coil 36 as indicated by match lines labeled 3 and 4. As long as switches 40 and 42 are in the open position as shown, no magnetic fields are generated by coils 34 and 36, and valves 30 and 32 will remain in a state of low thermal conductivity so long as their temperature is below the superconducting transition temperature. The thermal conductivity of valves 30 and 32 can be rendered normal by closing switches 40 and 42 to magnetically energize coils 34 and 36 respectively. The state of thermal conductivity of each of the valves 30 and 32 should be controllable independent of the conductive state of the other, as is the case in the present example.

To obtain electrocaloric refrigeration effects with Type I materials, it is necessary to polarize at least a portion of the polarizable agents of refrigerant 24 with an electric field, the existence of which field is represented by an arrow labeled E. To accomplish this, a pair of spaced plate or film electrodes 44 are disposed on opposite surfaces of the refrigerant 24 to form a capacitorlike assembly. Electrodes 44 are electrically connected by lead wires 46 which pass through a pair of feedthrough insulators 48, to an adjustable high voltage supply 50.

Supply 50 may be of any well-known type capable of providing a voltage across electrodes 44 and consequently, an electric field E through refrigerant 24 of sufficient intensity to at least partially polarize the latter. For example, the supply 50 may consist of a DC voltage source 52 connected by means of an enabling switch 54 across a suitable voltage dividing network 56. The voltage applied across the electrodes 44 is taken from the network 56 between one side thereof and an adjustable tap or contact 58. By adjusting contact 58, it is possible to

reduce the voltage across electrodes 44 from a maximum value equal to the voltage across source 52, to 0 volts.

In order to monitor the temperature of refrigerant 24 during the operating sequence, a thermometer 60, such as the well-known germanium resistance type, may be used. A lead wire 62 from thermometer 60 may be brought through pipe 33 and feedthrough insulator 64 where it may be connected to any suitable utilization device, not shown, for temperature display and control purposes. Alternately, the refrigerant may be used as its own thermometer, as for example, by measuring the capacitance across electrodes 44. The high-voltage supply could consist of a conventional control circuit 65 which provides an output voltage to maintain the refrigerant at temperature  $T_1$  in which case lead wire 62 would be connected directly to the voltage supply as illustrated in FIG. 1b.

Referring now to FIGS. 1 and 2, a typical Carnot-cycle operating sequence of the system of the present invention will be explained in connection with Type I materials. Although numerous other operating modes may be employed, the Carnot-cycle mode of operation is preferred since its use yields the highest possible thermodynamic efficiency for any selected set of operating temperatures.

Assume that the system of FIG. 1 is initially at room temperature with all the switches 40, 42, and 54 in an open position as shown. It will therefore be necessary to establish certain initial conditions within the system before the operating sequence can begin. First, a suitable starting temperature  $T_1$  must be selected which is within the electrocaloric refrigeration range of refrigerant 24. This selection will depend partly upon the physical nature of the particular electrocaloric material employed as refrigerant 24, and partly upon the ultimate low temperature  $T_2$  desired to be obtained. Although most useful Type I electrocaloric materials show some slight tendency to cool paraelectrically at relatively high temperatures ( $100^\circ\text{K.}$ ), it will probably be found necessary to start the process from a much lower temperature, preferably in the range of pumped liquid helium below  $1.5^\circ\text{K.}$ , where a substantial percentage temperature decrease can be obtained. For Type I materials the reduction in entropy by changing the polarization of such material assures that the temperature of the material will drop by a substantial percentage when the material is subsequently depolarized adiabatically. The Type I electrocaloric materials whose properties have been fairly thoroughly investigated include polar impurity doped alkali-halides such as NaCl:OH, KCl:OH, RbCl:CN, and the like and polar impurity doped rutilelike crystalline material such as  $\text{TiO}_2\text{:OH}$ ,  $\text{SbO}_2\text{:OH}$ , and the like.

Since the paraelectric properties of KCl:OH appear to be well understood in the art at the present time, this material has been selected for refrigerant 24 of the present example for the purpose of illustrating the use of Type I electrocaloric materials. However, it should be mentioned that any of the above-mentioned or other suitable Type I materials can be used successfully in the present invention. Provided they are doped with a polar impurity to within a reasonable range of concentrations, each will produce a substantial temperature decrease paraelectrically when the process is begun from a starting temperature  $T_1$  within the range of liquid helium.

A suitable starting temperature  $T_1$  having been selected, refrigerant 24, consisting of about 20 cc. of KCl containing about  $10^{18}$  OH ions/cc., load 26 and valves 30 and 32 are refrigerated by refrigerator 10 since a heat conductive path extends from helium 16 through lid 28, valve 30, refrigerant 24, and valve 32 to the load. It may be found desirable to pump a gas such as helium into chamber 22 to improve thermal contact between the articles within chamber 22 and helium 16 by way of thermally conductive lid 28 during this cooling step. Once the assembly within the chamber 22 is brought into thermal equilibrium with helium 16 at the desired starting temperature  $T_1$ , chamber 22 is evacuated, preferably to about 1 micron of Hg or less, in order to provide conditions wherein refrigerant 24 and load 26 can be effectively isolated thermally when required during the operating sequence to follow.

Valves 30 and 32, then being at an operative temperature, can be utilized as required to provide the necessary thermal connections between refrigerant 24, load 26, and helium 16.

With refrigerant 24 at the desired starting temperature  $T_1$ , switch 40 may be closed to provide thermal contact between helium 16 and the refrigerant under the evacuated condition of the chamber 22. Then switch 54 is closed and contact 58 is adjusted so as to apply a potential to electrodes 44 thereby generating a field  $E$ , preferably of sufficient intensity  $E_1$  to align substantially all of the polarizable agents contained in the Type I refrigerant 24, this condition being known as saturation polarization. The symbols  $E_1$ ,  $E_2$ ,  $E_3$ , and  $E_4$  are situated adjacent to equipotential lines passing through points 1, 2, 3, and 4, respectively. Though an intensity  $E_1$  of less than that necessary to produce saturation polarization may be used, either the operating temperature difference will be reduced or the period during which  $T_2$  can be maintained (line segment 2-3) will be reduced. Using the refrigerant of the present example, an electric field intensity  $E_1$  of about 60 kv./cm. will be sufficient to produce the desired state of saturation at about 1.3° K. in KCl:OH. A greater intensity  $E_1$  could be employed up to the dielectric breakdown strength of refrigerant 24, but this could not produce a higher level of polarization and therefore would not result in any greater operating advantage.

Immediately after the field  $E$  is applied, a heat of polarization is generated within refrigerant 24. However, since refrigerant 24 is thermally connected to helium 16 by valve 30, the heat of polarization will be absorbed by the helium and the polarization step will be brought about in a substantially isothermal manner. Accordingly, the entropy of refrigerant 24 will be substantially reduced below that of its unpolarized condition, to an entropy of  $S_1$ . The initial conditions  $S_1$ ,  $E_1$ , and  $T_1$  of the refrigerant 24, as represented by the point 1 in FIG. 2, are now established so that the operating sequence of the system can begin. Switch 40 should now be opened to thermally isolate refrigerant 24 from helium 16.

Refrigerant 24 is thereafter partially depolarized by adjusting contact 58 to reduce field  $E$  from intensity  $E_1$  to a lower intensity  $E_2$  (greater than 0 volts/cm.). In general, the lower the intensity  $E_2$ , the lower will be the final temperature  $T_2$  to which refrigerant 24 will cool, so long as refrigerant 24 remains within its electrocaloric refrigeration range. The low temperature limit attainable will be in the fraction of a degree Kelvin range where quantum effects known in the art as zero field splitting effects become appreciable.

It is advantageous to only partially depolarize refrigerant 24 rather than completely depolarize it, since by partial depolarization, as by adjusting contact 58, the final temperature  $T_2$  can be selected quite accurately. Also, refrigerant 24 can be employed at the heart of a Carnot-cycle refrigeration system because the residual intensity  $E_2$  remaining after partial depolarization can be advantageously employed to maintain isothermal conditions by offsetting the effects of heat leakage into refrigerant 24 from heat generated in the load, for example.

Refrigerant 24, having been partially depolarized, will cool adiabatically along the line segment 1-2 as thermal energy within refrigerant 24 is converted into potential energy during the disalignment of a quantity of the polarizable agents within the refrigerant material refrigerant material relative to the direction of the field  $E$ . Equilibrium conditions will be established and the condition of refrigerant 24 will stabilize at point 2. Starting from point 1 under saturation conditions of about 60 kv./cm. at a temperature  $T_1$  of about 1.3° K. it should be possible to cool refrigerant 24 by partial adiabatic depolarization to a final temperature  $T_2$  of about 0.1° K. with an electric field intensity  $E_2$  of about 8.0 kv./cm. remaining. However, some adjustment in the intensity  $E_2$  may be required in order to attain the precisely desired final temperature  $T_2$ .

Now refrigerant 24 is at the final temperature  $T_2$ , and assuming the first cycle of operation of the system is being executed, load 26 is still at the starting temperature  $T_1$ , since valve 32 is still closed. At this point in later successive cycles

of operation, however, the temperature of load 26 will approach and reach temperature  $T_2$ , after which time the operation of the refrigerator is to maintain the load at  $T_2$ .

To cool load 26, switch 42 is closed to thermally connect load 26 and refrigerant 24. Accordingly, refrigerant 24 begins absorbing heat from the warmer load thus tending to reduce the temperature of the latter. For maximum thermodynamic efficiency, it is important that refrigerant 24 absorb heat from load 26 in a substantially isothermal manner so that the entropy of refrigerant 24 will increase along the line segment 2-3 until thermal equilibrium is eventually established at the point 3. Clearly, however, without taking additional measures, the temperature of refrigerant 24 will tend to increase as it absorbs heat thus departing from the desired operating path to follow a curve of constant field intensity  $E_2$  as represented by the dashed curve 66. To correct this tendency, contact 58 is adjusted as soon as a slight temperature increase is detected by the thermometer 60 so as to reduce the field  $E$  by a sufficient increment to return refrigerant 24 to the desired temperature  $T_2$ . The dashed line segment 68 represents the slight decrease in temperature of refrigerant 24 brought about by this incremental or partial depolarization. From time-to-time further reductions in the field  $E$  should be made to offset the continuing effect of heat absorption by refrigerant 24. This can readily be accomplished either manually by observing the temperature of refrigerant 24, or automatically through the use of any well-known temperature responsive automatic control means. In either case the line segment 2-3 can be adhered to quite closely so that an excellent quality Carnot-cycle sequence can be approached. Eventually the field will be reduced to zero and point 3 will be reached. In the time period that elapses during the portion of the cycle represented by line segment 2-3, the refrigerant absorbs heat from the load, and this condition is represented by the arrows labeled  $Q_{23}$ .

Ideally there should have been just enough field intensity at point 2 to have been completely expended along the line segment 2-3 such that the field  $E_3$  at point 3 is zero. Though this would be desirable in improving the refrigeration capacity of the system, it is not essential to the success of the process. Therefore, some quantity of field  $E$  could exist at the point 3 as represented by  $E_3$ .

The load 26 is then thermally isolated from refrigerant 24 by opening switch 42, and contact 58 is adjusted to increase the intensity of the field to a value  $E_4$  just sufficient to increase the temperature of refrigerant 24 to the starting temperature  $T_1$ . In effect refrigerant 24 is now being repolarized so that its temperature increases as the heat of polarization is generated. The step is adiabatic since it will be observed that refrigerant 24 is in a thermally isolated condition at this time, both the valves 30 and 32 being in a nonconductive condition.

When the field has been increased to a value  $E_4$  so that the temperature of refrigerant 24 is at  $T_1$ , refrigerant 24 is thermally connected to helium 16 by closing switch 40 and energizing valve 30 to a state of high thermal conductivity. Since both refrigerant 24 and helium 16 are at the same temperature  $T_1$  no heat transfer will take place as a result. However, now contact 58 is further adjusted to increase the intensity of field  $E$  from  $E_4$  to the saturation intensity  $E_1$ . Since refrigerant 24 is now thermally connected to helium 16 the heat of polarization is absorbed by the latter and the step occurs isothermally at the starting temperature  $T_1$  along line segment 4-1. The arrows labeled  $Q_{41}$  point away from line segment 4-1 to represent the heat rejected by the refrigerant. Equilibrium conditions are established at point 1, the initial cycle of operation of the system being completed, and a similar new cycle can now begin.

It should be noted that where the thermal insulation of the system is high, and where the load does not generate heat, it may not be necessary to begin a new refrigeration cycle immediately. The maximum permissible time lapse between the end of a cycle and the beginning of a succeeding cycle will depend on the rate of heat leakage into the system and the rate of heat generated by the load, if any. Consequently, the time

lapse between cycles depends on how rapidly load 26 tends to increase in temperature from the desired value  $T_2$ . If the load were a source of heat, the system could be operated in a continuous series of cycles, one immediately following another, as is commonly the case with many well-known closed-cycle refrigerators.

The refrigerator of FIGS. 1 and 1a can operate in a cycle different from that described above when Type II materials are utilized. However, reference may still be made to FIG. 2 for a description of this modified operating cycle, since only the valves of the electric field intensities  $E_1$ ,  $E_2$ ,  $E_3$  and  $E_4$  are changed. A family of electrocaloric materials has been discovered that functions as Type II materials, i.e., the materials cool when adiabatically polarized. These materials are strontium titanate-containing glasses from which the strontium titanate, for example perovskite strontium titanate, has been controllably crystallized, and results in a body composed of a multiplicity of uniform very fine grain crystals substantially homogeneously dispersed in a glassy matrix and forming a large proportion thereof. Such materials can be obtained when at least 50 percent by weight of strontium titanate, including 0 to 5 percent by weight of dopants, is melted into a glass at a temperature of about 1,650° C., and a low-temperature ferroelectric strontium titanate can be subsequently crystallized in this glass upon reheating between 800° and 1,300° C., for a time period of at least one hour at 800° C. and at least ½ minute at 1,300° C. A more complete discussion of a suitable heat treatment process used for the formation of crystals may be found in the A. Herczog and S. D. Stookey patent, U.S. Pat. No. 3,195,030 which is incorporated herein by reference. With proper heat treatment, strontium titanate is the major crystallized phase that will result. It is to be understood that said glass can be heat treated prior to use in or formation of an ultimate device, or such device may be formed using the glass and thereafter be subjected to heat treatment whereby the strontium titanate is crystallized in situ. Small additions of up to 5 percent by weight may be made of oxides such as niobium oxide, tantalum oxide, vanadium oxide and bismuth oxide. Such dopants tend to stabilize the titanate phase. Suitable glasses for forming the glass matrix of the present invention may be alumino-silicate, borate, phosphate, or germanite glasses.

Table I represents examples of compositions on a batch basis in percent by weight which, when heat treated as herein set forth, provide material comprising strontium titanate crystals in a glass matrix suitable for the purposes of the present invention.

TABLE I

Ex.	SrTiO <sub>3</sub>	SrCO <sub>3</sub>	TiO <sub>2</sub>	SiO <sub>2</sub>	Al <sub>2</sub> O <sub>3</sub>	Nb <sub>2</sub> O <sub>5</sub>	V <sub>2</sub> O <sub>5</sub>	Ta <sub>2</sub> O <sub>5</sub>	Bi <sub>2</sub> O <sub>3</sub>
1.		40.4	26.7	10.9	4.0				
2.	72.0			18.7	9.3				
3.	87.8	12.5		21.1	8.6				
4.	68.4	0.6		20.0	10.0	1.0			
5.	62.8	2.7		19.8	9.9	4.8			
6.		47.8	25.1	21.5	4.3	1.7			
7.	68.4	0.6		20.0	10.0		1.0		
8.	62.8	2.7		19.8	9.9		4.8		
9.		47.4	25.1	21.5	4.3		1.7		
10.	68.4	0.6		20.0	10.0			1.0	
11.	62.8	2.7		19.8	9.9			4.8	
12.		47.4	25.1	21.5	4.3			1.7	
13.	68.4	0.6		20.0	10.0				1.0
14.	62.8	2.7		19.8	9.9				4.6
15.		47.4	25.1	21.5	4.3				1.7
16.		50.4	27.3	18.6	3.7				
17.		50.1	27.1	19.0	3.8				
18.		47.9	25.9	17.5	8.7				
19.		49.1	26.6	19.4	4.9				

With a Type II material refrigerant incorporated in the refrigerator of FIGS. 1 and 1a, the operation of the refrigerator is similar to that described in connection with Type I refrigerants, except as follows. Initial conditions are established at point 1 of FIG. 2 by reducing the temperature of refrigerant 24 and load 26 to temperature  $T_1$  as described above, but the entropy of refrigerant 24 is reduced by reducing the electric field therein to zero. If this is the initial cycle of operation the zero field state in which the Type II material

naturally exists will merely be maintained. With both valves 30 and 32 closed so that no heat flows to or from the refrigerant, the electric field applied to the Type II refrigerant is increased from its zero value to an intermediate value, causing the temperature of the refrigerant to drop from temperature  $T_1$  to a preselected lower temperature  $T_2$ . This condition is represented by point 2 on the curve of FIG. 2. Valve 32 is then opened so that heat can flow from the load to the refrigerant. As heat is absorbed by the refrigerant the temperature thereof increases and the temperature-entropy relationship starts to follow curve 66. Since maximum efficiency is obtained when the operating cycle is caused to follow line segment 2-3, the field intensity is increased slightly as soon as a slight temperature increase is detected by thermometer 60 so that the temperature of refrigerant 24 is again reduced to the desired temperature  $T_2$ . The dashed line segment 68 represents the slight decrease in temperature of refrigerant 24 brought about by this incremental polarization resulting from the increased electrical field E. Further increases in the field are made as necessary to offset any tendency of the refrigerant to sustain a temperature increase as it absorbs heat from the load represented by arrows  $Q_{23}$ . As in the previous description, the temperature-entropy relationship can be caused to follow line segment 2-3 by observing or sensing the temperature of the refrigerant as measured by thermometer 60 and then appropriately increasing the voltage applied to electrodes 44. The refrigerant can thereby be maintained at a temperature  $T_2$  until saturation polarization is reached at a maximum electric field  $E_3$ . In FIG. 2, this condition is represented by point 3. An increase in field intensity above saturation intensity  $E_3$  cannot produce any further entropy change.

Switch 42 is now opened to thermally isolate load 26 from refrigerant 24. With both valves 30 and 32 closed, the field is decreased to a value  $E_4$  to warm the refrigerant to the first-stage temperature  $T_1$ . This adiabatic temperature increase follows line segment 3-4. Valve 30 is then opened so that heat can flow between the refrigerant and the first-stage, and the electric field on the refrigerant is isothermally decreased to zero. During this step heat, represented by arrows  $Q_{41}$ , flows from the refrigerant to the first-stage and the temperature of the refrigerant is maintained constant at  $T_1$ . Equilibrium conditions are thereby again established at point 1 and another cycle can be started.

Although the present invention has been described with respect to specific details of certain embodiments thereof, it is not intended that such details be limitations on the present invention. Although maximum efficiency is obtained by causing

the apparatus of the present invention to operate on a Carnot-cycle, this invention encompasses any closed-cycle type of operation. For example, after the refrigerant cools to temperature  $T_2$  at point 2 of FIG. 2, and load 26 is connected thereto, the temperature of the refrigerant could be permitted to increase along dashed line 66 until temperature  $T_1$  is reached. Thereafter, field  $E_1$  could be established and the refrigerant could be connected to the first stage by valve 30 to bring the system back to the initial condition at point 1. Another modifi-

cation in the operating cycle could be achieved by simultaneously applying field  $E_1$  to the refrigerant while opening valve 30 after the operating cycle has progressed along line segments 1-2 and 2-3 of FIG. 2. The result would be to move from point 3 to point 1. It is obvious that there are many other modes of operation too numerous to mention, but which are within the scope of the following claims.

I claim:

1. A closed-cycle process for refrigerating a mass of material comprising the steps of providing an electrocaloric refrigerant, thermally connecting said refrigerant and material to a sink having a temperature of  $T_1$  so that said refrigerant and material are substantially cooled to temperature  $T_1$ , thermally isolating said material, isothermally changing the polarization of said refrigerant toward a first state of polarization by applying thereto an electric field of intensity  $E_1$  to reduce the entropy of said refrigerant to a value  $S_1$ , the condition of said refrigerant wherein the entropy thereof is at value  $S_1$ , the temperature thereof is at value  $T_1$ , and the electric field applied thereto is of value  $E_1$  constituting an initial condition, thermally isolating said refrigerant, adiabatically changing the polarization of said refrigerant to a second state of polarization by changing the intensity of said field from said intensity  $E_1$  to an intensity  $E_2$  to reduce the temperature of said refrigerant from said temperature  $T_1$  to a temperature  $T_2$ , and thermally connecting said refrigerant to said material to decrease the temperature of said material.

2. The process of claim 1 wherein, after the step of thermally connecting said refrigerant to said material, said initial condition is reestablished.

3. The process of claim 2 wherein said step of reestablishing said initial condition comprises changing said field from said intensity  $E_2$  to an intensity  $E_3$  such that said refrigerant absorbs heat from said material substantially isothermally and the entropy of said refrigerant increases from said value  $S_1$  to a value  $S_2$ , said change in field intensity from  $E_2$  to  $E_3$  being in the same direction as the change in field intensity from  $E_1$  to  $E_2$ .

4. The process of claim 3 wherein said step of reestablishing said initial conditions further comprises

adiabatically changing said field intensity to an intensity  $E_4$ , said change in field intensity from  $E_3$  to  $E_4$  being in the direction opposite to the change in field intensity from  $E_1$  to  $E_2$ , to increase the temperature of said refrigerant from said temperature  $T_2$  to said temperature  $T_2$ , and isothermally changing said field from said intensity  $E_4$  to said intensity  $E_1$  to reduce the entropy of said refrigerant from said value  $S_2$  to said value  $S_1$ .

5. The process of claim 4 wherein said field intensity  $E_1$  is sufficiently high to substantially produce saturation polarization of said refrigerant and said field intensity  $E_3$  is substantially 0.

6. The process of claim 4 wherein said field intensity  $E_1$  is substantially 0.

7. The process of claim 1 wherein said refrigerant is a material selected from the group consisting of electrically polar ion doped alkali-halides, electrically polar ion doped rutilelike crystalline materials, and glass ceramics consisting of a glass matrix, 0 to 5 percent by weight of at least one dopant, and at least 50 percent by weight of strontium titanate, wherein at least a portion of said strontium titanate is crystallized.

8. A closed-cycle refrigerator for refrigerating a mass of material, said refrigerator comprising an electrocaloric refrigerant, a sink for absorbing heat from said refrigerant at a substantially constant temperature within the electrocaloric refrigeration range of said refrigerant, means for thermally connecting and separating said

refrigerant and said material, means for thermally connecting and separating said refrigerant and sink,

means for thermally isolating said refrigerant and material when said last recited means thermally separate said refrigerant and sink, and

means for electrically polarizing and depolarizing said refrigerant.

9. The refrigerator of claim 8 wherein said refrigerant consists essentially of a host crystalline material containing a polar ion dopant selected from the group consisting of OH and CN ions.

10. The refrigerator of claim 8 wherein said refrigerant consists of a material selected from the group consisting of electrically polar ion doped alkali-halides, electrically polar ion doped rutilelike crystalline materials, and glass ceramics consisting of a glass matrix, 0 to 5 percent by weight of at least one dopant, and at least 50 percent by weight of strontium titanate, wherein at least a portion of said strontium titanate is crystallized.

11. The refrigerator of claim 8 wherein said sink comprises a helium cryogenic refrigerator.

12. The refrigerator of claim 8 wherein said means for thermally connecting and separating said refrigerant and sink comprises a first magnetically actuated thermal valve connected between said refrigerant and sink.

13. The refrigerator of claim 8 wherein said means for thermally connecting and separating said refrigerant and material comprises a second magnetically actuated thermal valve connected between said refrigerant and material.

14. The refrigerator of claim 8 wherein said means for thermally isolating comprises a vacuum chamber in which said refrigerant and material are disposed.

15. The refrigerator of claim 8 wherein said means for polarizing and depolarizing said refrigerant comprises a variable high-voltage source, and a pair of electrodes separated by said refrigerant, said electrodes being connected across said source.

16. The refrigerator of claim 8 further comprising means for detecting the temperature of said refrigerant, and means responsively connected to said detecting means for changing the polarization of said refrigerant in response to a temperature rise therein, for maintaining said refrigerant at a substantially constant temperature while said refrigerant absorbs heat from said material.

17. The refrigerator of claim 12 wherein said means for thermally connecting and separating said refrigerant and material comprises a second magnetically actuated thermal valve connected between said refrigerant and material, and wherein said means for polarizing and depolarizing said refrigerant comprises a high-voltage source and a pair of electrodes separated by said refrigerant, said electrodes being connected across said high-voltage source.

18. The refrigerator of claim 17 wherein said means for thermally isolating comprises a vacuum chamber in which said refrigerant and material are disposed.

19. The refrigerator of claim 18 wherein said refrigerant consists essentially of a host crystalline material containing a polar ion dopant selected from the group consisting of OH and CN ions.

20. The refrigerator of claim 19 wherein said sink comprises a helium cryogenic refrigerator.

21. The refrigerator of claim 18 wherein said refrigerant consists of a material selected from the group consisting of electrically polar ion doped alkali-halides, electrically polar ion doped rutilelike crystalline materials, and glass ceramics consisting of a glass matrix, 0 to 5 percent by weight of at least one dopant, and at least 50 percent by weight of strontium titanate, wherein at least a portion of said strontium titanate is crystallized.

22. The refrigerator of claim 21 wherein said sink comprises a helium cryogenic refrigerator.

\* \* \* \* \*

UNITED STATES PATENT OFFICE  
CERTIFICATE OF CORRECTION

Patent No. 3,638,440 Dated February 1, 1972

Inventor(s) William N. Lawless

It is certified that error appears in the above-identified patent and that said Letters Patent are hereby corrected as shown below:

Column 7, Table 1, Example 6, "47.8" should be

-- 47.4 --; Example 14, "4.6" should be

-- 4.8 --.

Claim 4, line 7, " $T_2$ " (second occurrence) should be

--  $T_1$  --.

Signed and sealed this 1st day of August 1972.

(SEAL)  
Attest:

EDWARD M. FLETCHER, JR.  
Attesting Officer

ROBERT GOTTSCHALK  
Commissioner of Patents